

# C2 - THE PERIODIC TABLE 2.6 The transition metals

## GCSE Chemistry Aiming for Grade 4

Where are the transition metals on the periodic table?  
In the centre block

Give the names of five transition metals

**Gold, Silver, Aluminium, Iron, Copper**

List the physical properties of the transition metals

1. Good conductors of heat
2. Good conductors of electricity
3. High densities
4. High melting points
5. Hard and strong

Why is mercury not typical as a transition metal?

It is a liquid at room temperature  
whereas all the others are solid with  
high melting points.



Assessment  
Grade 4

## GCSE Chemistry Aiming for Grade 6

Compared to Group 1 metals how do the following  
properties compare?

**Density**

Much higher than group 1, none of them will float on  
water like lithium, sodium and potassium.

**Melting point**

Much higher melting points except for mercury

**Boiling point**

Boiling points are extremely high compared to the  
alkali metals

**Hardness**

Alkali metals can be cut with a knife and are soft.  
Transition metals are very hard.

**Strength**

Transition metals are very strong and can be used for  
buildings and machines due to their strength

Give the names of the following transition metal  
compounds.

**$\text{Fe}_2\text{O}_3$**

Iron Oxide

**$\text{Cu}_2\text{O}$**

Copper Oxide

**$\text{NiCO}_3$**

Nickel Carbonate

**$\text{AgNO}_3$**

Silver Nitrate

Assessment  
Grade 6

## GCSE Chemistry Aiming for Grade 8

Why is copper used for wires?  
It is a very good conductor -  
better than most other metals. It is also very  
durable and does not rust.

Steel is made from Iron and  
carbon. It is used for bridges?  
It is very strong and hard to  
break and has a high melting point.

Titanium is used for aircraft  
Titanium has a very high melting point  
and is very strong and hard to break.

State the colours of the compounds  
of transition metals.

Copper sulphate blue due to

Nickel Carbonate pale green

Chromium oxide dark green

Manganese Chloride pale pink

Compared with Group 1 elements  
metals have different properties  
Transition metals can form a  
variety of ions. Their reactivity is different.  
eg Iron can form  $\text{Fe}^{2+}$  and  $\text{Fe}^{3+}$   
Different reactivity.

Assessment  
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