

10. Electrolysis of copper – refining copper

Electrodes: metallic copper

Electrolyte: copper sulfate solution

Ions Present: Cu^{2+} , SO_4^{2-} , OH^- , H^+

Half equations:

In the Cathode $\text{Cu}^{2+} + 2e^- \rightarrow \text{Cu}$

In the Anode $\text{Cu} \rightarrow \text{Cu}^{2+} + 2e^-$

Overall Equation $\text{Cu}^{2+} + \text{Cu} \rightarrow \text{Cu} + \text{Cu}^{2+}$

Observations:

Cathode: Dark solid deposit (pure copper)

Anode: Disintegration of anode and yellow deposit below

Uses of copper: electrical wiring, copper plating

Why is it necessary to purify Copper? In this way, it will have more conductivity

What are the elements present in the slag? Ag, Au, Pt



(show in the graph how it will look like)

11. Electroplating with Copper

Electrodes:

Cathode: Spoon (metal more reactive than Cu)

Anode: Pure copper

Electrolyte: Copper Sulfate or Nitrate

Ions Present: Cu^{2+} , SO_4^{2-} , OH^- , H^+

Half equations:

In the Cathode $\text{Cu}^{2+} + 2e^- \rightarrow \text{Cu}$

In the Anode $\text{Cu} \rightarrow \text{Cu}^{2+} + 2e^-$

Overall Equation $\text{Cu}^{2+} + 2e^- + \text{Cu} \rightarrow \text{Cu} + \text{Cu}^{2+}$

Observations:

Cathode: the color of the spoon becomes darker

Anode: Disintegrate Copper



Matter And Change Electrochemistry Study Guide Answer

Richard Watts



Matter And Change Electrochemistry Study Guide Answer:

Chemistry, Student Study Guide James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH_4 from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

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