

Question 83
Acids and Bases IV
Advanced

10. A buffer is 0.10 M acetic acid (HAc) and 0.10 M acetate ion (Ac⁻). What is the concentration of H⁺ in this buffer? Assume that the total ionic strength is 0.10 M.

$$K_a(\text{HAc}) = 1.8 \times 10^{-5} \text{ M}$$

$$\text{HAc} \rightleftharpoons \text{H}^+ + \text{Ac}^-$$

What is the pH of this buffer? The addition of H⁺ ions will shift the equilibrium of the dissociation of HAc to the left, according to Le Chatelier's principle. Because K_a is small, we can assume that [H⁺] is small. The pH of the buffer is 4.74. The pH of the buffer is 4.74. The pH of the buffer is 4.74. The pH of the buffer is 4.74.

11. A buffer is 0.10 M acetic acid (HAc) and 0.10 M acetate ion (Ac⁻). What is the pH of this buffer? Assume that the total ionic strength is 0.10 M.

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Lecture 32 Acids And Bases Iv Answers

A Gutmann



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