

4.7 CHEMISTRY (13%)

4.7.1 Chemistry Paper 1 (13%)

4. (a) Ether (1 mark)
- (b) Add water to each compound.
Add Na_2CO_3 or NaHCO_3 to a suitable part of each sample.
 $\text{C}_2\text{H}_5\text{CO}$ no effervescence (1 mark)
 $\text{C}_2\text{H}_5\text{CO}$ effervescence (1 mark)
on add acidified potassium dichromate (VII) acid reacts
 $\text{C}_2\text{H}_5\text{CO}$ colourless change to green
 $\text{C}_2\text{H}_5\text{CO}$ no change
on add acidified Potassium Manganate (VII)
 $\text{C}_2\text{H}_5\text{CO}$ no decolourised
 $\text{C}_2\text{H}_5\text{CO}$ no change (1 mark)
5. (a) Ether (NaOH) (1 mark)
- (b) — Sodium is very reactive (more electropositive) (1 mark)
— More reactive than carbon
6. (a) Ether
Sodium lamp, used in molecular reactions
Sodium vapour, sodium amalgams
(NaOH), Esterification and saponification, etc. (1 mark)
7. (a) Endothermic change, when one mole of crystal lattice is broken into its ions in gaseous state. (1 mark)
- (b) Endothermic reaction (percent) (1 mark)
8. (a) Boyle's law: The volume of a fixed mass of a gas is inversely proportional to its pressure or the product of pressure and volume is constant at a fixed temperature. (1 mark)

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The Enigmatic Realm of **January 2015 Chemistry Igcse Mark Scheme**: Unleashing the Language is Inner Magic

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