

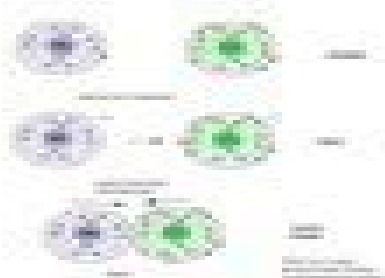
Final Study Guide

Valence Electrons: Electrons on the outermost shell or orbital, electrons involved in bonding, atoms are most stable when they have 8 (this is called the octet rule)

Isotopes: atoms with the same atomic number, but different atomic masses, because they have different amount of neutrons

Oxidation Numbers: positive or negative number shows the symbol that tells us how many electrons an atom has to lose or gain

Ionic Bond: a bond between a metal and a non-metal where electrons are lost and gained



Metal usually loses the electrons and the nonmetal usually gains it, this causes the metal to be positively charged and the nonmetal to be negatively charged. Positives and negatives attract.

Properties:

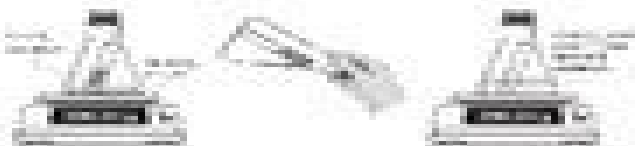
- Conduct electricity when melted or dissolved in water.
- High melting points
- Are polar
- Usually form crystalline structures

Covalent Bonds: a bond between two nonmetals where electrons are shared, both nuclei are attracted to the shared electrons. Electrons can be shared, unequally which results in a polar covalent bond. Can also form double and triple bonds.

Properties:

- Low melting points
- Tend to be softer
- Usually don't dissolve in water
- Don't conduct electricity

The Law of Conservation of mass: Mass cannot be created or destroyed. The mass of the products must equal the mass of the reactants



as you can see in this picture the mass before the reaction is equal to the mass after the reaction.

This rule applies to atoms as well.

You must have the same number and type of atoms at the beginning of the reaction as you do after the reaction

Steps for Balancing Equations

1. List the atoms involved in the reaction
2. Count how many of each type of atom on each side
3. Add coefficients to make the number of each type of atom on each side be equal
4. Recount the number of each type of atom on each side
5. If not equal repeat steps 3-4

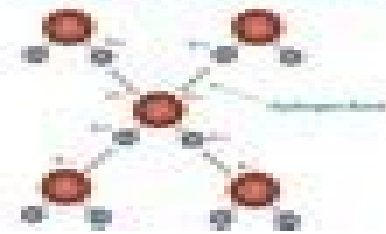
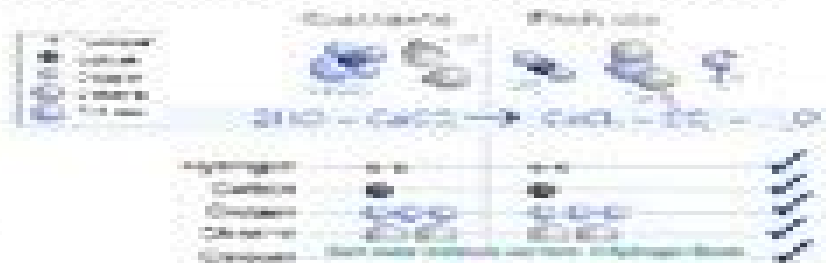
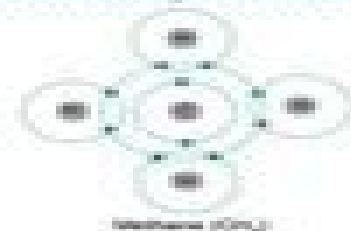
Structure of water molecule



Water

- Is a polar molecule

Electron sharing in covalent bond



IPC Study Guide

Carl Oskar



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diagnostic tests and treatments that are used to optimize care in the critically ill is essential. Since our first edition 7 years ago, new information crucial to the care and understanding of the critically ill patient has rapidly accumulated. Because this knowledge base crosses many different disciplines, a comprehensive multidisciplinary approach presenting the information is essential, similar to the multidisciplinary approach that is used to care for the critically ill patient. We have strived to provide this content in an easily digestible format that uses a variety of teaching tools to facilitate understanding of the presented concepts and to enhance information retention. To meet the demand to provide comprehensive and diverse education in order to understand the pathogenesis and optimum care of a variety of critical illnesses, we have substantially revised the prior topics in the first edition with updated information. We have also markedly expanded the number of topics covered to include acute lung injury and the acute respiratory distress syndrome, an expanded discussion of the physiology and operation of mechanical ventilation, obstetrical care in the ICU, neurosurgical emergencies, acute coronary syndromes, cardiac arrhythmias, role of whole body rehabilitation in the ICU, ethical conduct of human research in the ICU, and nursing care of the ICU patient.

The Guide to Interpersonal Psychotherapy Myrna M. Weissman, John C. Markowitz, Gerald L. Klerman, 2017-08-01

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