

# Ionic Chemical Formulas/PreAICE Chemistry

NAME: \_\_\_\_\_ PERIOD \_\_\_\_\_

**Directions:** Using the Chemical Formulas program sheet, fill in the table below. Use subscripts & superscripts. Remember to use roman numerals to show the valence of the transition metals in the compound name.

METAL CATION (+)	CATION FORMULA	NONMETAL ANION (-)	ANION FORMULA	COMPOUND FORMULA	COMPOUND NAME
sodium	Na <sup>+1</sup>	hydroxide	OH <sup>-</sup>	NaOH	sodium hydroxide
magnesium	Mg <sup>+2</sup>	phosphate	PO <sub>4</sub> <sup>-3</sup>	Mg <sub>3</sub> (PO <sub>4</sub> ) <sub>2</sub>	magnesium phosphate
calcium	Ca <sup>+2</sup>	chloride	Cl <sup>-</sup>	CaCl <sub>2</sub>	calcium chloride
gold(I)	Au <sup>+1</sup>	sulfite	SO <sub>3</sub> <sup>-2</sup>	Au <sub>2</sub> SO <sub>3</sub>	gold(I) sulfite
copper(II)	Cu <sup>+2</sup>	hydroxide	OH <sup>-</sup>	Cu(OH) <sub>2</sub>	copper(II) hydroxide
Iron(III)	Fe <sup>+3</sup>	telluride	Te <sup>-2</sup>	Fe <sub>2</sub> Te <sub>3</sub>	iron(III) telluride
magnesium	Mg <sup>+2</sup>	bromide	Br <sup>-1</sup>	MgBr <sub>2</sub>	magnesium bromide
manganese(II)	Mn <sup>+2</sup>	sulfide	S <sup>-2</sup>	MnS	manganese(II) sulfide
boron	B <sup>+3</sup>	oxide	O <sup>-2</sup>	B <sub>2</sub> O <sub>3</sub>	boron oxide
potassium	K <sup>+</sup>	nitrate	NO <sub>3</sub> <sup>-</sup>	KNO <sub>3</sub>	potassium nitrate
barium	Ba <sup>+2</sup>	bromide	Br <sup>-</sup>	BaBr <sub>2</sub>	barium bromide
strontium	Sr <sup>+2</sup>	sulfate	SO <sub>4</sub> <sup>-2</sup>	SrSO <sub>4</sub>	strontium sulfate
iron(III)	Fe <sup>+3</sup>	oxide	O <sup>-2</sup>	Fe <sub>2</sub> O <sub>3</sub>	iron(III) oxide
silver(I)	Ag <sup>+1</sup>	chloride	Cl <sup>-</sup>	AgCl	silver(I) chloride
aluminum	Al <sup>+3</sup>	carbonate	CO <sub>3</sub> <sup>-2</sup>	Al <sub>2</sub> (CO <sub>3</sub> ) <sub>3</sub>	aluminum carbonate
gold(II)	Au <sup>+2</sup>	nitride	N <sup>-3</sup>	Au <sub>3</sub> N <sub>2</sub>	gold(II) nitride
cesium	Cs <sup>+1</sup>	phosphate	PO <sub>4</sub> <sup>-3</sup>	Cs <sub>3</sub> PO <sub>4</sub>	cesium phosphate
ammonium	NH <sub>4</sub> <sup>+</sup>	sulfide	S <sup>-2</sup>	(NH <sub>4</sub> ) <sub>2</sub> S	ammonium sulfide
zinc(II)	Zn <sup>+2</sup>	nitrite	NO <sub>2</sub> <sup>-</sup>	Zn(NO <sub>2</sub> ) <sub>2</sub>	zinc(II) nitrite
chromium(I)	Cr <sup>+1</sup>	nitride	N <sup>-3</sup>	Cr <sub>3</sub> N	chromium(I) nitride
mercury(II)	Hg <sup>+2</sup>	oxide	O <sup>-2</sup>	HgO	mercury(II) oxide
sodium	Na <sup>+</sup>	sulfite	SO <sub>3</sub> <sup>-2</sup>	Na <sub>2</sub> SO <sub>3</sub>	sodium sulfite
magnesium	Mg <sup>+2</sup>	oxide	O <sup>-2</sup>	MgO	magnesium oxide

# Ions Answer Key

**Effiong Eyo**

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the name bar.

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dependent on the charge in contrast to the behavior previously found for mixed solutions of  $\text{NH}_4\text{NO}_3$  and also  $\text{NH}_4\text{ClO}_4$  with  $\text{NH}_4\text{F}$  at constant ionic strength where the dependence is approximately quadratic

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