

Ch 5 Ions and Ionic Compounds Study Guide**Multiple Choice**

Identify the choice that best completes the statement or answers the question.

- _____ 1. Which of the following electron configurations belongs to an element that is NOT chemically reactive?
a. $1s^2 2s^2 2p^3$ c. $1s^2 2s^2 2p^5$
b. $1s^2 2s^2 2p^6 3s^2 3p^4$ d. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^1$
- _____ 2. In the compound sodium fluoride, NaF, the sodium atom loses one electron and the fluorine atom gains one electron to form ions that have electron configurations similar to
a. helium. c. neon.
b. oxygen. d. calcium.
- _____ 3. In many compounds, atoms of main-group elements form ions so that the number of electrons in the outermost energy levels of each ion is
a. 2. c. 8.
b. 6. d. 10.
- _____ 4. The electron configuration of nitrogen is $1s^2 2s^2 2p^3$. How many more electrons does nitrogen need to have an electron configuration similar to neon?
a. 1 c. 5
b. 3 d. 8
- _____ 5. Atoms of copper and iron
a. generally form stable bonds with transition elements.
b. have stable electron configurations.
c. tend to form cations.
d. tend to form anions.
- _____ 6. An anion
a. is an ion with a negative charge.
b. attracts ions with negative charges.
c. results when an alkaline-earth metal loses one of its two outermost electrons.
d. has more protons than electrons.
- _____ 7. When the octet rule is satisfied, the outermost _____ are filled.
a. *d* and *f* orbitals c. *s* and *d* orbitals
b. *s* and *p* orbitals d. *d* and *p* orbitals
- _____ 8. The elements of the _____ group satisfy the octet rule without forming compounds.
a. halide c. alkali metal
b. noble gas d. alkaline-earth metal
- _____ 9. Once an atom has full *s* and *p* orbitals in its outermost energy level,
a. it is highly reactive only with alkali metals.
b. it is highly reactive only with halogens.
c. it can be combined with most elements.
d. it has a stable octet and is unreactive.

Ions And Ionic Compounds Study Guide

Arshad Iqbal



Ions And Ionic Compounds Study Guide:

Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03

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Chemistry, Student Study Guide James E. Brady, Fred Senese, 2008-01-28. The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst. Carbon nanotubes are a recently discovered allotrope of carbon. Three other allotropes of carbon, buckyballs, graphite, and diamond, are illustrated at the left, as is the molecule methane, CH_4 , from which nanotubes and buckyballs can be made. The element carbon forms an amazing number of compounds with structures that follow from simple methane, found in natural gas, to the complex macromolecules that serve as the basis of life on our planet. The study of chemistry also follows from the simple to the more complex, and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement.

The Practice of Chemistry Study Guide & Solutions Manual Pamela Mills, Amina El-Ashmawy, 2003-04-14. Designed to help students understand the material better and avoid common mistakes. Also includes solutions and explanations to odd-numbered exercises.

Chemical Principles Study Guide/Solutions Manual John Krenos, Joseph Potenza, Carl Hoeger, 2007-01-18. Written for general chemistry courses, Chemical Principles helps students develop chemical insight by showing the connection between chemical principles and their applications.

Study Guide and Solutions Manual for Fundamentals of General, Organic, and Biological Chemistry Susan McMurry, 1992.

Study Guide to Accompany Calculus for the Management, Life, and Social Sciences Clyde Metz, 1984-01-01. Study Guide to Accompany Calculus for the Management, Life, and Social Sciences.

Student Study Guide/Solutions Manual for Essentials of General, Organic, and Biochemistry Denise Guinn, Rebecca Brewer, Rachel C. Lum, 2009-09-15. The Student Study Guide and Solutions Manual provides students with a combined manual designed to help them avoid common mistakes and understand key concepts. After a brief review of each section's critical ideas, students are taken through stepped-out worked examples, try-it-yourself examples, and chapter quizzes.

all structured to reinforce chapter objectives and build problem solving techniques The solutions manual includes detailed solutions to all odd numbered exercises in the text

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Chemical Principles Student's Study Guide & Solutions Manual John Krenos, Joseph Potenza, Loretta Jones, Lynn Koplitz, Thomas Spence, 2004-03-19 This combination manual is designed to help students avoid common mistakes and understand the material better The solutions manual section includes detailed answers and explanations to the odd numbered exercises in the text

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