

Questions

2. Following the rules for stabilisation for the last two questions, write correct chemical formulae for each of the following compounds:

Compound	Formula
(a) Lithium chloride	$\text{Li}^{+1} / \text{Cl}^{-1} \rightarrow \text{LiCl}$
(b) Magnesium nitride	$\text{Mg}^{+2} / \text{N}^{-3} \rightarrow \text{Mg}_3\text{N}_2$
(c) Strontium sulfide	$\text{Sr}^{+2} / \text{S}^{-2} \rightarrow \text{SrS}$
(d) Rubidium fluoride	$\text{Rb}^{+1} / \text{F}^{-1} \rightarrow \text{RbF}$
(e) Lithium oxide	$\text{Li}^{+1} / \text{O}^{-2} \rightarrow \text{Li}_2\text{O}$
(f) Sodium selenide	$\text{Na}^{+1} / \text{Se}^{-2} \rightarrow \text{Na}_2\text{Se}$
(g) Potassium chloride	$\text{K}^{+1} / \text{Cl}^{-1} \rightarrow \text{KCl}$
(h) Calcium phosphide	$\text{Ca}^{+2} / \text{P}^{-3} \rightarrow \text{Ca}_3\text{P}_2$
(i) Barium oxide	$\text{Ba}^{+2} / \text{O}^{-2} \rightarrow \text{BaO}$
(j) Aluminium nitride	$\text{Al}^{+3} / \text{N}^{-3} \rightarrow \text{AlN}$

3. Use your answers to the Key Questions and the Exercise Questions to draw a conclusion about the order of ions in each compound of the elements in the compounds you drew the most groups (strontium, rubidium and aluminium nitride, lithium chloride and potassium chloride).

Ionic Bonding Chemistry If8766 Answers

Margaret Irene Lindsay



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