

Ionic Compounds and Metals

Section 7.1 Ion Formation

pages 206–209

Section 7.1 Assessment

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1. **Compare** the stability of a lithium atom with that of its ion, Li^+ .

The Li^+ ion is more stable because it has a complete octet.

2. **Describe** two different causes of the force of attraction in a chemical bond.

The two causes are the attraction between the positive nucleus of one atom and the negative electrons of another atom, and the attraction between positive ions and negative ions.

3. **Apply** Why are all of the elements from group 18 relatively unreactive, whereas those in group 17 are very reactive?

The group 18 elements, known as noble gases, have complete outer energy levels and do not easily form ions. The group 17 elements are highly reactive because they need only to gain a single electron to form an octet.

4. **Summarize** ionic bond formation by correctly pairing these terms: *cation*, *anion*, *electron gain*, *electron loss*.

(anion, electron gain), (cation, electron loss)

5. **Apply** Write out the electron configuration for each atom. Then predict the change that must occur in each to achieve a noble gas configuration.

- a. nitrogen

$[\text{He}]2s^22p^3$; gain 3 electrons (3^- ion) or lose 5 electrons (5^+ ion)

- b. sulfur

$[\text{Ne}]3s^23p^4$; gain 2 electrons (2^- ion)

- c. barium

$[\text{Xe}]6s^2$; lose 2 electrons (2^+ ion)

- d. lithium

$[\text{He}]2s^1$; lose 1 electron (1^+ ion)

6. **Model** Draw models to represent the formation of the positive calcium ion and the negative bromide ion.

Models should show that the calcium atom loses two electrons, forming Ca^{2+} , and that bromine gains one electron, forming Br^- . The models should also show the addition of energy to form the calcium ion, and the release of energy to form the bromide ion.

Section 7.2 Ionic Bonds and Ionic Compounds

pages 210–217

Practice Problems

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Explain how an ionic compound forms from these elements.

7. sodium and nitrogen

Three Na atoms each lose $1 e^-$, forming 1^+ ions. One N atom gains $3 e^-$, forming a 3^- ion. The ions attract, forming Na_3N .

$$3 \text{ Na ions } \left(\frac{1+}{\text{Na ion}} \right) + 1 \text{ N ion } \left(\frac{3-}{\text{N ion}} \right) \\ = 3(1+) + 1(3-) = 0$$

The overall charge on one formula unit of Na_3N is zero.

8. lithium and oxygen

Two Li atoms each lose $1 e^-$, forming 1^+ ions. One O atom gains $2 e^-$, forming a 2^- ion. The ions attract, forming Li_2O .

$$2 \text{ Li ions } \left(\frac{1+}{\text{Li ion}} \right) + 1 \text{ O ion } \left(\frac{2-}{\text{O ion}} \right) \\ = 2(1+) + 1(2-) = 0$$

The overall charge on one formula unit of Li_2O is zero.

Ionic Compounds And Metals Solutions Manual

**Philip C. Keller, Raymond Chang, Jill L.
Keller**



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The Practice of Chemistry Study Guide & Solutions Manual Pamela Mills, Amina El-Ashmawy, 2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises

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Fundamentals of Environmental Sampling and Analysis Chunlong Zhang, 2024-07-03 This is the Solutions Manual to accompany Fundamentals of Environmental Sampling and Analysis Second Edition It provides solutions to the exercises and problems found in the main volume This book introduces a comprehensive overview on the fundamentals and applications of environmental sampling and analysis for students in environmental science and engineering as well as environmental professionals involved in sampling and analytical work The book details fundamentals of sampling selection of standard methods QA QC sample preparation chemical and instrumental principles and method applications to various contaminants in environmental matrices air water soil waste and biological samples The book gives an integrated introduction to sampling and analysis both are essential to quality environmental data For example contrary to other books that introduce a specific area of sampling and analysis this text provides a balanced mix of field sampling and laboratory analysis essential knowledge in chemistry statistics hydrology regulations wet chemical methods for conventional chemicals as well as various modern instrumental techniques for contaminants of emerging concerns The new edition adds three standalone chapters regarding the basics of analytical and organic chemistry environmental data analysis mass spectrometry and other significant amounts of new materials such as time integrated passive sampling incremental sampling green sample preparation Raman spectroscopy chiral separation and non target analysis In addition the second edition provides more examples visual aids case studies and end of chapter exercise problems to enhance a better understanding of the fundamentals of environmental sampling and analysis while incorporating current literature mostly peer reviewed journal papers regarding the applications and challenges in the field of

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