F325 Past paper Answers

- (a) G: CO (1) HCOOH/H₂SO₄ → CO + H₂O (1)
 - (b) H: C(1) $C_{12}H_{22}O_{11} \rightarrow 12C + 11H_2O(1)$ 2

2

3

3

5

1

[7]

[13]

.

(c) I $C_4H_6O_2$ (1) $2C_2H_6O_2 \rightarrow C_4H_6O_2 + 2H_2O$ (1) Structure:

accept any sensible structure of C₄H₄O₂

- (a) empirical formula
 N: O = 63.64/14: 36.36/16 (1)
 = 4.56: 2.27 = 2: 1. Empirical formula = N₂O (1)
 molecular formula
 A4, of gas = 1.833 × 24 = 44 (1)(calc 43.992)
 with these two pieces of evidence, assume that molecular formula = N₂O
 - (b) any chemical that reacts to produce gas:
 e.g. carbonate and CO₂ (1)
 accept: metal more reactive than Pb and H₂
 balanced equation to match chemical added (1)
 - (c) Af₂(Lidocaine) = 236 (1) moles Novocaine = 100 × 10⁻³/236 = 4.24 × 10⁻⁴ (1) concentration of Novocaine = 4.24 × 10⁻⁴ × (1000/5) = 0.0347/0.0848 mol dm⁻³ (1)
 - (d) mass C = 12 × 3.74/44.0 = 12 × 0.085 = 1.02 g (1) mass H = 2/18 × 0.918 = 0.102 g (1) mass O = 1.394 - (1.020 + 0.102) = 0.272 g ratio C : H : O = 1.02/12 : 0.102/1 : 0.272/16 (1) = 0.0850 : 0.102 : 0.0170 / 5 : 6 : 1 / C₃H₆O (1) C₃H₁₀O has relative mass of 82 M₆ is 164 = 2 × 82 ∴ molecular formula = C₁₀H₁₂O₂ (1)
- (i) 4PCl₅ + 10MgO → P₄O₁₀ + 10MgCl₂ (a)
 - (ii) $100g P_4O_{10} = 284 / 0.35(2) \text{ mol (1)}$ $moles PCI_5 \text{ needed} = 4 \times 0.352 = 1.408/1.4 \text{ mol (1)}$ $mass PCI_5 = 1.4(08) \times 208.5 = 293.568 / 294 g/ 291.9 g (1)$

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