

INDORE PUBLIC SCHOOL
MID TERM EXAM PRACTICE PAPER

Class-XI

SUBJECT – CHEMISTRY (THEORY)

Time : 3 Hrs

M.M. : 70

General instructions: -

- (i) This question paper contains 26 questions. Write answers neatly and legibly. All questions are compulsory. This question paper has five sections: Section A, Section B, Section C, Section D and Section E.
- (ii) Question no. 1 to 5 are very short answer questions and carry 1 mark each.
- (iii) Question no. 6 to 10 are short answer questions and carry 2 marks each.
- (iv) Question no. 11 to 22 are also short answer questions and carry 3 marks each.
- (v) Question no. 23 is value based question and carry 4 marks.
- (vi) Question no. 24 to 26 are long answer questions and carry 5 marks each.
- (vii) Use log tables, if necessary. Use of calculators is not allowed.

SECTION-A [5 Questions (1 Mark Each), Total Marks: 5]

- Q.1** How many numbers of atoms are present in 1.4 gram of nitrogen gas?
- Q.2** Write the electronic configuration of Chromium and Copper? (At. No. are 24, 29 respectively)
- Q.3** What is the n_{factor} of KMnO_4 in acidic medium?
- Q.4** Define ionization energy?
- Q.5** Give hybridisation of XeO_2F_2 and CO_3^{2-} .

SECTION-B [5 Questions (2 Marks Each), Total Marks: 10]

- Q.6** Nitrogen and Hydrogen gets combined to give Ammonia 0.5 mole of N_2 and 2.0 moles of H_2 are taken initially. Find:
 - i) What is limiting reagent?
 - ii) Weight of NH_3 gets formed?
- Q.7** Define Aufbau rule and Pauli's Exclusion Principle.
- Q.8** What are the reasons behind diagonal relationship and write the elements which shows the diagonal relationship?
- Q.9** Among F and Cl, which one have the higher electron affinity. Give reason for your answer.
- Q.10** Which is more basic, NH_3 or PH_3 ? Give reason for your answer.

SECTION-C [12 Questions (3 Marks Each), Total Marks: 36]

- Q.11**
 - a) Define Heisenberg Uncertainty Principle?
 - b) Minimum number of electron required to form a bond are?
- Q.12** Give the hybridisation and structure of PCl_3F_2 .
- Q.13** The magnetic moment of an element is 3.8 B.M. ($Z = 27$).
 - a) No. of unpaired electrons are.
 - b) Write electronic configuration of the element.
- Q.14** A compound contains C = 71.23%, H = 12.95% and O = 15.81%. What is the empirical formula of the compound?
- Q.15** Define Molarity (M), Molality (m) and Normality (N).

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