

Ionic Bonding Worksheet

For each pair of elements below draw an atomic diagram showing electrons in different energy levels. Draw arrows to show where the outer electrons will go during a chemical reaction, then draw the resulting compound. Finally, fill in the table below each reaction. Refer to the sample shown.



Atoms	Valence electrons	Electron transfer from/to each atom	Ions formed in the product
Li	1 ve	transfers 1ve <sup>-</sup> to F	$\text{Li}^{+}$
F	7 ve	accepts 1ve <sup>-</sup> from Li	$\text{F}^{-}$

Reactions	Atoms	Valence electrons	Electron transfer from/to each atom	Ions formed in the product
1) $\text{Li} + \text{Cl} \Rightarrow \text{LiCl}$ $\text{Li} + \cdot\ddot{\text{Cl}}: \rightarrow [\text{Li}]^{+1} [\ddot{\text{Cl}}:]^{+1}$	Li	1 ve	donate 1e <sup>-</sup>	$\text{Li}^{+1}$
	Cl	1 ve	accepts 1e <sup>-</sup>	$\text{Cl}^{-1}$
2) $\text{Ca} + \text{O} \Rightarrow \text{CaO}$ $\text{Ca} + \cdot\ddot{\text{O}}: \rightarrow [\text{Ca}]^{+2} [\ddot{\text{O}}:]^{2-}$	Ca	2 ve	donates 2e <sup>-</sup>	$\text{Ca}^{+2}$
	O	6 ve	accepts 2e <sup>-</sup>	$\text{O}^{-2}$
3) $\text{Be} + \text{F} \Rightarrow \text{BeF}_2$ $\text{Be} + \cdot\ddot{\text{F}}: \rightarrow [\text{Be}]^{+2} [\ddot{\text{F}}:]^{1-}$	Be	2 ve	donates 2e <sup>-</sup> to each F	$\text{Be}^{+2}$
	F	7 ve	each F accepts 1ve each	$\text{F}^{-1}$
4) $\text{Mg} + \text{S} \Rightarrow \text{MgS}$ $\text{Mg} + \cdot\ddot{\text{S}}: \rightarrow [\text{Mg}]^{+2} [\ddot{\text{S}}:]^{2-}$	Mg	2 ve	donates 2e <sup>-</sup> to S	$\text{Mg}^{+2}$
	S	6 ve	accepts 2e <sup>-</sup>	$\text{S}^{-2}$
5) $\text{K} + \text{F} \Rightarrow \text{KF}$ $\text{K} + \cdot\ddot{\text{F}}: \rightarrow [\text{K}]^{+1} [\ddot{\text{F}}:]^{1-}$	K	1 ve	donates 1ve <sup>-</sup>	$\text{K}^{+1}$
	F	7 ve	accepts 1ve <sup>-</sup>	$\text{F}^{-1}$

# Ionic Bonding Worksheet 1 Answer Key

**Sir Tarun Rupani**



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