

Radical reactions: Free-radical polymerisation

What is a radical? A polymerisation process initiated by radicals is called a free-radical polymerisation. Free-radical polymerisation is initiated by radicals, which are species with an unpaired electron. These species are formed by the homolysis of a molecule, which is a process in which a molecule is broken into two fragments, each with an unpaired electron. This process is called homolysis. The fragments are called radicals. The process of homolysis is initiated by heat or light. The fragments are called radicals because they have an unpaired electron. The process of homolysis is initiated by heat or light. The fragments are called radicals because they have an unpaired electron.

Radical reactions: Free-radical polymerisation **Free-radical polymerisation** is a process in which a monomer is converted into a polymer by the action of a radical. The process is initiated by a radical, which is a species with an unpaired electron. The radical reacts with the monomer to form a new radical, which then reacts with another monomer to form a polymer chain. The process continues until the radical is terminated by a species that has no unpaired electrons.

Radical reactions: Free-radical polymerisation			
Radical reactions: Free-radical polymerisation			
$\text{C}_2\text{H}_5\text{Br}$	ethyl bromide	Br^\bullet	ethyl bromide radical
HBr	hydrogen bromide	H^\bullet	hydrogen bromide radical
$\text{C}_2\text{H}_5\text{I}$	ethyl iodide	I^\bullet	ethyl iodide radical
$\text{C}_2\text{H}_5\text{Cl}$	ethyl chloride	Cl^\bullet	ethyl chloride radical
$\text{C}_2\text{H}_5\text{F}$	ethyl fluoride	F^\bullet	ethyl fluoride radical
$\text{C}_2\text{H}_5\text{OH}$	ethyl alcohol	OH^\bullet	ethyl alcohol radical
$\text{C}_2\text{H}_5\text{SH}$	ethyl thiol	SH^\bullet	ethyl thiol radical
$\text{C}_2\text{H}_5\text{NH}_2$	ethyl amine	NH_2^\bullet	ethyl amine radical
$\text{C}_2\text{H}_5\text{N}_3$	ethyl azide	N_3^\bullet	ethyl azide radical
$\text{C}_2\text{H}_5\text{N}_3\text{H}$	ethyl azide hydrazide	$\text{N}_3\text{H}^\bullet$	ethyl azide hydrazide radical
$\text{C}_2\text{H}_5\text{N}_3\text{H}_2$	ethyl azide hydrazide hydrazide	$\text{N}_3\text{H}_2^\bullet$	ethyl azide hydrazide hydrazide radical
$\text{C}_2\text{H}_5\text{N}_3\text{H}_3$	ethyl azide hydrazide hydrazide hydrazide	$\text{N}_3\text{H}_3^\bullet$	ethyl azide hydrazide hydrazide hydrazide radical

Radical reactions: Free-radical polymerisation

1. What is a radical? A radical is a species with an unpaired electron. Radicals are formed by the homolysis of a molecule, which is a process in which a molecule is broken into two fragments, each with an unpaired electron. The fragments are called radicals. The process of homolysis is initiated by heat or light. The fragments are called radicals because they have an unpaired electron.
2. What is a free-radical polymerisation? A free-radical polymerisation is a process in which a monomer is converted into a polymer by the action of a radical. The process is initiated by a radical, which is a species with an unpaired electron. The radical reacts with the monomer to form a new radical, which then reacts with another monomer to form a polymer chain. The process continues until the radical is terminated by a species that has no unpaired electrons.
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Ionic Bonding Advanced Chemquest 20 Polyatomic Ions

Anthony S. Fauci



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