#### Question 1

- (a) Using only the periodic table, arrange each set of atoms in order of increasing radius:
  - (1) Ba, Ca, Na

Na ≺ Ca ≺ Ba

(2) In. Sn. As

As < Sn < In

(3) Cs, K, Al.

Al < K < Cs

- (b) Which neutral atom is isoelectronic with each of the following ions?
  - (1) H\*

None

(The H\* ion has 0 electrons. It is impossible for a neutral atom to have 0 electrons.)

(2) Zr\*\*

POR.

(3) Mn<sup>7</sup>\*

Ar

(4) 11

Xie

(c) Explain the following variations in atomic or ionic radii:

(1)  $1^{-} > 1 > 1^{+}$ 

As the atomic number stays constant and the number of electrons increases, the electron-electron repulsions increase. This leads to the electrons spreading out, and the anion becomes larger. The reverse is also true, so the cation is smaller.

(2)  $Ca^{2*} > Mg^{2*} > Be^{2*}$ 

For cations with the same charge, ionic radii increase going down a column because of the increase in principal quantum number, which means the outer electrons are farther away from the nucleus on average. This leads to larger ion size.

# **Ionic Radius Answer Key**

**Disha Experts** 



This book delves into Ionic Radius Answer Key. Ionic Radius Answer Key is a vital topic that must be grasped by everyone, ranging from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Ionic Radius Answer Key, encompassing both the fundamentals and more intricate discussions.

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