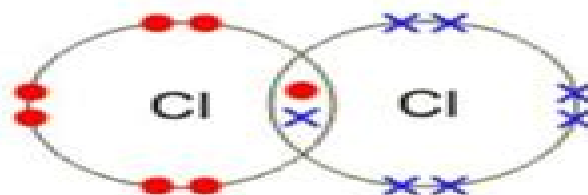


Covalent bonding

A non-metal metal and another non-metal can bond together by **sharing** electrons so that both electrons can achieve a **full outer shell**. This is called covalent bonding. To represent the atoms involved in the bonding we draw crosses for the electrons on one of the atoms and dots for the electrons on the other atoms as shown here.



Instructions

NB: It may be easiest to draw each atom out individually first to work out how many electrons will be shared

1. Draw the circles (outer shells) and overlap them (like a Venn diagram)
2. Write the symbols of the atoms involved in the middle of each circle
3. Draw the **electrons** for one atom as dots (o) and electrons for the other atom as crosses (x).
4. Any shared electrons are drawn in the **overlapping** section.
5. Count up the electrons for **EACH** atom and check that they each now have a **full outer shell** (2 electrons for the first shell and 8 for the next 2 shells)

TASK: Draw dot and cross diagrams for the following covalent compounds.

F ₂		CH ₄	
HCl		CO ₂	
H ₂ S		H ₂ O	
HF		NH ₃	

Challenge - O₂

Ionic And Metallic Bonding Questions

JG Myers



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Table of Contents Ionic And Metallic Bonding Questions

1. Understanding the eBook Ionic And Metallic Bonding Questions
 - The Rise of Digital Reading Ionic And Metallic Bonding Questions
 - Advantages of eBooks Over Traditional Books
2. Identifying Ionic And Metallic Bonding Questions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ionic And Metallic Bonding Questions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ionic And Metallic Bonding Questions
 - Personalized Recommendations
 - Ionic And Metallic Bonding Questions User Reviews and Ratings

- Ionic And Metallic Bonding Questions and Bestseller Lists
- 5. Accessing Ionic And Metallic Bonding Questions Free and Paid eBooks
 - Ionic And Metallic Bonding Questions Public Domain eBooks
 - Ionic And Metallic Bonding Questions eBook Subscription Services
 - Ionic And Metallic Bonding Questions Budget-Friendly Options
- 6. Navigating Ionic And Metallic Bonding Questions eBook Formats
 - ePub, PDF, MOBI, and More
 - Ionic And Metallic Bonding Questions Compatibility with Devices
 - Ionic And Metallic Bonding Questions Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ionic And Metallic Bonding Questions
 - Highlighting and Note-Taking Ionic And Metallic Bonding Questions
 - Interactive Elements Ionic And Metallic Bonding Questions
- 8. Staying Engaged with Ionic And Metallic Bonding Questions
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ionic And Metallic Bonding Questions
- 9. Balancing eBooks and Physical Books Ionic And Metallic Bonding Questions
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ionic And Metallic Bonding Questions
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Ionic And Metallic Bonding Questions
 - Setting Reading Goals Ionic And Metallic Bonding Questions
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ionic And Metallic Bonding Questions
 - Fact-Checking eBook Content of Ionic And Metallic Bonding Questions
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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