

Percent Composition and Molecular Formula Worksheet

1. Calculate the percent of nitrogen in urea, NH_2CONH_2 .
2. Calculate the percentage of water in zinc sulfate heptahydrate.
3. Calculate the percentage of oxygen in potassium chlorate.
4. Calculate the percentage of tin in potassium stannate trihydrate, $\text{K}_2\text{SnO}_3 \cdot 2\text{H}_2\text{O}$

Write the molecular formulas of the following compounds:

5. A compound with an empirical formula of C_2OH_4 and a molar mass of 88 grams per mole.
6. A compound with an empirical formula of $\text{C}_4\text{H}_4\text{O}$ and a molar mass of 136 grams per mole.
7. A compound with an empirical formula of CFBrO and a molar mass of 254.7 grams per mole.
8. A compound with an empirical formula of $\text{C}_2\text{H}_5\text{N}$ and a molar mass of 46 grams per mole.

Answer the following questions:

9. The percentage composition of acetic acid is found to be 39.9% C, 6.7% H, and 53.4% O. Determine the empirical formula of acetic acid.
10. The molar mass for question #9 was determined by experiment to be 60.0 g/mol. What is the molecular formula?
11. What's the empirical formula of a molecule containing 18.7% lithium, 16.3% carbon, and 65.0% oxygen?
12. A 50.51 g sample of a compound made from phosphorus and chlorine is decomposed. Analysis of the products showed that 11.39 g of phosphorus atoms were produced. What is the empirical formula of the compound?
13. When 2.5000 g of an oxide of mercury, (Hg_xO_y) is decomposed into the elements by heating, 2.405 g of mercury are produced. Calculate the empirical formula.
14. The compound benzamide has the following percent composition. What is the empirical formula?
C = 69.40 % H= 5.825 % O = 13.21 % N= 11.57 %
15. A component of protein called serine has an approximate molar mass of 100 g/mole. If the percent composition is as follows, what is the empirical and molecular formula of serine?
C = 34.95 % H= 6.844 % O = 46.56 % N= 13.59 %
16. Aniline, a starting material for urethane plastic foams, consists of C, H, and N. Combustion of such compounds yields CO_2 , H_2O , and N_2 as products. If the combustion of 9.71 g of aniline yields 6.63 g H_2O and 1.46 g N_2 , what is its empirical formula?
17. The molar mass of aniline is 93 g/mol. What is its molecular formula?

Formula Writing Cavalcade

SJ Ball



Formula Writing Cavalcade:

Reviewing **Formula Writing Cavalcade**: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is really astonishing. Within the pages of "**Formula Writing Cavalcade**," an enthralling opus penned by a very acclaimed wordsmith, readers attempt an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve in to the book is central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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