

Chemistry Formulas



Dilution

$$M_1V_1 = M_2V_2$$

Energy of a photon

$$E_{\text{photon}} = hf = hc/\lambda$$

Pressure

$$P = F/A$$

Average Kinetic Energy

$$KE_{\text{avg}} = \frac{3}{2}RT$$

Ideal Gas Law

$$PV = nRT$$

Combined Gas Law

$$\frac{P_1V_1}{n_1T_1} = \frac{P_2V_2}{n_2T_2}$$

Dalton's Law of Partial Pressures

$$P_{\text{total}} = P_A + P_B + \dots$$

Dalton's Law of Partial Pressures

$$P_A = \chi_A P_{\text{total}}$$

Graham's Law of Effusion

$$\frac{r_1}{r_2} = \sqrt{\frac{M_2}{M_1}}$$

Henry's Law

$$P_A = k_H[A]$$

Freezing Point Depression

$$\Delta T_F = -iK_F M$$

Boiling Point Elevation

$$\Delta T_B = iK_B m$$

Vapor Pressure Depression (Raoult's Law)

$$P_{\text{soln}} = \chi_{\text{soln}} P_{\text{soln}}^0$$

Osmotic Pressure (π)

$$\pi = iMRT$$

Arrhenius Equation

$$k = Ae^{-E_a/RT}$$

Equilibrium Constant Expressions

$$K_c = \frac{[\text{products}]}{[\text{reactants}]}$$

Equilibrium Constant Expressions

$$K_{eq} = \frac{k_{\text{forward}}}{k_{\text{reverse}}}$$

Equilibrium Constant Expressions

$$K_P = \frac{P_{\text{products}}}{P_{\text{reactants}}}$$

Reaction Quotient (Q)

$$Q = \frac{[\text{products}]}{[\text{reactants}]}$$

Solubility Product Constant (K_{sp})

$$K_{sp} = \frac{[\text{products}]}{[\text{reactants}]}$$

pH & pOH

$$pH = -\log[H^+]$$

pH & pOH

$$pOH = -\log[OH^-]$$

pH & pOH

$$pH + pOH = 14$$

pK_a & pK_b

pK_a & pK_b

pK_a & pK_b

Neutralization Reaction

Buffers

Enthalpy of Formation

First Law of Thermodynamics

Pressure Volume Work

Calorimetry Thermal Energy (q)

Heat Curves & Thermal Energy (q)

Heat Curves & Thermal Energy (q)

Heat Curves & Thermal Energy (q)

Entropy (S)

Bond Dissociation Energy

Gibb's Free Energy (ΔG)

Gibb's Free Energy (ΔG)

Gibb's Free Energy (ΔG)

Standard Cell Potential

Standard Cell Potential

Nernst Equation

Kinetics (always 1st order)

Kinetics (always 1st order)

Kinetics (always 1st order)

Nuclear Binding Energy

$$pK_a = -\log[K_a]$$

$$pK_b = -\log[K_b]$$

$$pK_a + pK_b = 14$$

$$n_a M_a V_a = n_b M_b V_b$$

$$pH = pK_a + \log[A^-]/[HA]$$

$$\Delta H_f = \sum \Delta H_f^{\circ}(\text{product}) - \sum \Delta H_f^{\circ}(\text{reactants})$$

$$\Delta E = q + w$$

$$w = -P\Delta V$$

$$q = -C_{\text{calorimeter}}\Delta T$$

$$q = mc\Delta T$$

$$q = m\Delta H_{\text{fusion}}$$

$$q = m\Delta H_{\text{vaporization}}$$

$$\Delta S = \ln S_{\text{products}} - \ln S_{\text{reactants}}$$

$$\Delta H = \sum \Delta H_{\text{reactants}} - \sum \Delta H_{\text{products}}$$

$$\Delta G^{\circ} = \Delta H^{\circ} - T\Delta S^{\circ}$$

$$\Delta G = \Delta G^{\circ} + RT\ln Q$$

$$\Delta G^{\circ} = -RT^{\circ}\ln(K_{eq})$$

$$E^{\circ} = E^{\circ}_{\text{reduction}} + E^{\circ}_{\text{oxidation}}$$

$$E^{\circ} = E^{\circ}_{\text{cathode}} + E^{\circ}_{\text{anode}}$$

$$E_{\text{cell}} = E^{\circ} - \frac{0.0592}{n} \log Q$$

$$N = N_0 e^{-kt}$$

$$\ln N = \ln N_0 - kt$$

$$t_{1/2} = \frac{0.693}{k}$$

$$E = \Delta mc^2$$

Formula Sheet For Gen Chem Chm 2045

Emilie Sanchez



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Table of Contents Formula Sheet For Gen Chem Chm 2045

1. Understanding the eBook Formula Sheet For Gen Chem Chm 2045
 - The Rise of Digital Reading Formula Sheet For Gen Chem Chm 2045
 - Advantages of eBooks Over Traditional Books
2. Identifying Formula Sheet For Gen Chem Chm 2045
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Formula Sheet For Gen Chem Chm 2045
 - User-Friendly Interface
4. Exploring eBook Recommendations from Formula Sheet For Gen Chem Chm 2045
 - Personalized Recommendations
 - Formula Sheet For Gen Chem Chm 2045 User Reviews and Ratings
 - Formula Sheet For Gen Chem Chm 2045 and Bestseller Lists
5. Accessing Formula Sheet For Gen Chem Chm 2045 Free and Paid eBooks
 - Formula Sheet For Gen Chem Chm 2045 Public Domain eBooks
 - Formula Sheet For Gen Chem Chm 2045 eBook Subscription Services
 - Formula Sheet For Gen Chem Chm 2045 Budget-Friendly Options
6. Navigating Formula Sheet For Gen Chem Chm 2045 eBook Formats

- ePub, PDF, MOBI, and More
- Formula Sheet For Gen Chem Chm 2045 Compatibility with Devices
- Formula Sheet For Gen Chem Chm 2045 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Formula Sheet For Gen Chem Chm 2045
 - Highlighting and Note-Taking Formula Sheet For Gen Chem Chm 2045
 - Interactive Elements Formula Sheet For Gen Chem Chm 2045
- 8. Staying Engaged with Formula Sheet For Gen Chem Chm 2045
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Formula Sheet For Gen Chem Chm 2045
- 9. Balancing eBooks and Physical Books Formula Sheet For Gen Chem Chm 2045
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Formula Sheet For Gen Chem Chm 2045
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Formula Sheet For Gen Chem Chm 2045
 - Setting Reading Goals Formula Sheet For Gen Chem Chm 2045
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Formula Sheet For Gen Chem Chm 2045
 - Fact-Checking eBook Content of Formula Sheet For Gen Chem Chm 2045
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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