

NATURAL SCIENCES

Gr. 9 - LESSON PLAN - MATTER & MATERIALS

TOPIC: COMPOUNDS		
Sub-topic: Names of Compounds		Duration: 1 Hour
CONTENT & CONCEPTS (CAPS p63)		
KEY CONCEPTS 1. Chemical reaction 2. Chemical formula 3. Names of compounds according to their elements.		
Specific Aims:	Specific Aim 1: 'Doing Science'	X
	Specific Aim 2: 'Knowing the subject content and making connections'	X
	Specific Aim 3: 'Understanding the uses of Science'	
LESSON OBJECTIVES: 1. Learners will be able to write the names of compounds as indicated on page 63 of the CAPS Document. 2. Learners will be able to make models of different compounds		
RESOURCES REQUIRED: • Periodic Table Poster • Siyavula Sasol Inzalo (Learner book) page 163		
TEACHING & LEARNING ACTIVITY: Introduction <u>Briefly review and revise concepts dealt with in Grade 8, focusing on compounds</u> As each element has a unique name for example: - Hydrogen - Lithium Remember the first 20 elements learnt in grade 8. How did you memorize them? Do you remember Mnemonics... He - H; He - He; Lies - Li; Because - Be; Boys - B; Can - C; Not - N; Operate - O; Fireplaces - F; New - Ne; Nation - Na; Might - Mg; Also - Al; Sign - Si; Peace - P; Security - S; Clause - Cl A - Ar; King - K; Can - Ca LEARNERS MAY RECITE THIS MNEMONIC AS A WAY OF MEMORIZING THE FIRST 20 ELEMENTS. <u>The teacher provides the following explanation to the learners</u> Each compound has a unique name. How is a compound formed? - A compound is formed when TWO or more different elements combine chemically. - Each atom in a compound are held by chemical bonds - Compounds are formed during chemical reaction (when elements/ compounds combine to form new substances) - Each compound can be written as a formula, for example, H₂O; NaCl; CuSO₄ What is a formula ? - This is a notation that shows which elements the compound is made up of. - It indicates the total number of atoms in one unit of the compound - It indicates the ratio in which atoms of different elements are bonded to each other - It gives the total number of atoms in one unit of a compound. For example, the formula for Copper Sulphate is CuSO ₄		

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Roman Wölfel



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Table of Contents Lesson Plans For Natural Science Grade 9

1. Understanding the eBook Lesson Plans For Natural Science Grade 9
 - The Rise of Digital Reading Lesson Plans For Natural Science Grade 9
 - Advantages of eBooks Over Traditional Books
2. Identifying Lesson Plans For Natural Science Grade 9
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Lesson Plans For Natural Science Grade 9
 - User-Friendly Interface
4. Exploring eBook Recommendations from Lesson Plans For Natural Science Grade 9
 - Personalized Recommendations
 - Lesson Plans For Natural Science Grade 9 User Reviews and Ratings
 - Lesson Plans For Natural Science Grade 9 and Bestseller Lists
5. Accessing Lesson Plans For Natural Science Grade 9 Free and Paid eBooks
 - Lesson Plans For Natural Science Grade 9 Public Domain eBooks
 - Lesson Plans For Natural Science Grade 9 eBook Subscription Services

- Lesson Plans For Natural Science Grade 9 Budget-Friendly Options
- 6. Navigating Lesson Plans For Natural Science Grade 9 eBook Formats
 - ePub, PDF, MOBI, and More
 - Lesson Plans For Natural Science Grade 9 Compatibility with Devices
 - Lesson Plans For Natural Science Grade 9 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Lesson Plans For Natural Science Grade 9
 - Highlighting and Note-Taking Lesson Plans For Natural Science Grade 9
 - Interactive Elements Lesson Plans For Natural Science Grade 9
- 8. Staying Engaged with Lesson Plans For Natural Science Grade 9
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Lesson Plans For Natural Science Grade 9
- 9. Balancing eBooks and Physical Books Lesson Plans For Natural Science Grade 9
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Lesson Plans For Natural Science Grade 9
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Lesson Plans For Natural Science Grade 9
 - Setting Reading Goals Lesson Plans For Natural Science Grade 9
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Lesson Plans For Natural Science Grade 9
 - Fact-Checking eBook Content of Lesson Plans For Natural Science Grade 9
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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