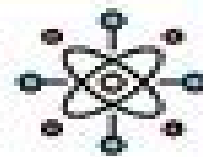


CHEMICAL BONDS



IONIC, COVALENT, & METALLIC

FORMS OF CHEMICAL BONDS
Bonding: combination of atoms

Atoms combine to form molecules by sharing or transferring electrons. This process is called bonding. The resulting molecule is more stable than the individual atoms.

Types of Chemical Bonds:

- 1. Ionic Bond:** Formed by the transfer of electrons from one atom to another, resulting in oppositely charged ions that attract each other.
- 2. Covalent Bond:** Formed by the sharing of electron pairs between two atoms.
- 3. Metallic Bond:** Formed by the attraction between positive metal ions and a sea of delocalized electrons.

COVALENT BONDING AND MOLECULAR STRUCTURE

Covalent bonding occurs when two atoms share one or more pairs of electrons. This sharing results in a stable molecule. The structure of a molecule is determined by the arrangement of its atoms and the type of bonds between them.

Diagram: A diagram showing the Lewis structure of water (H_2O). The oxygen atom is in the center, bonded to two hydrogen atoms. Each bond represents a pair of shared electrons. The oxygen atom has two lone pairs of electrons.

Periodic Table: A small periodic table is shown at the bottom, highlighting the groups of elements that typically form ionic and covalent bonds.



IONIC BONDING

Ionic bonding occurs when one atom transfers an electron to another atom, resulting in the formation of oppositely charged ions. These ions are then attracted to each other, forming an ionic bond.

Diagram: A diagram showing the formation of an ionic bond between a sodium atom (Na) and a chlorine atom (Cl). The sodium atom loses an electron to become a sodium ion (Na^+), and the chlorine atom gains an electron to become a chloride ion (Cl^-). The resulting ions are shown with their respective charges and are attracted to each other.

COVALENT BONDING

Covalent bonding occurs when two atoms share one or more pairs of electrons. This sharing results in a stable molecule. The structure of a molecule is determined by the arrangement of its atoms and the type of bonds between them.

Diagram: A diagram showing the Lewis structure of a diatomic molecule, such as oxygen (O_2). The two oxygen atoms are shown with their valence electrons. A double bond is formed by the sharing of four electrons (two pairs) between the two atoms. Each oxygen atom has two lone pairs of electrons.

MOLECULAR STRUCTURE

The shape of a molecule is determined by the arrangement of its atoms and the type of bonds between them. This is known as molecular structure. The shape of a molecule can be predicted using the VSEPR model.

Diagram: A diagram showing the VSEPR model for a water molecule (H_2O). The oxygen atom is in the center, bonded to two hydrogen atoms. The lone pairs of electrons on the oxygen atom are shown as two pairs of dots. The resulting molecular shape is bent.

GUIDED READING ACTIVITY

Ionic Bonds Guided Reading And Study Answers

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