

This question is about a compound called **sodium chloride**.

- (a) How many elements are joined together to form sodium chloride? **2** (1)
 (b) Write the chemical formula of sodium chloride. **NaCl** (1)
 (c) Name the type of bonding in sodium chloride. **ionic** (1)
 (d) Complete the sentence.

Ions are atoms, or groups of atoms, which have lost or gained **electrons** (1)

- (e) Suggest **three** properties which sodium chloride has.

Property 1 **Hard / crystalline** (1)

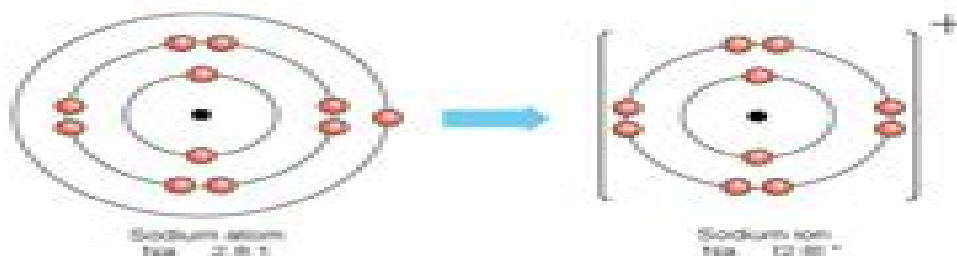
Property 2 **High melting / boiling point** (1)

Property 3 **Soluble / conducts electricity when in solution / molten** (3)

- (f) Write a balanced symbol equation for the reaction between magnesium (Na) and chlorine (Cl₂) to form sodium chloride (NaCl).

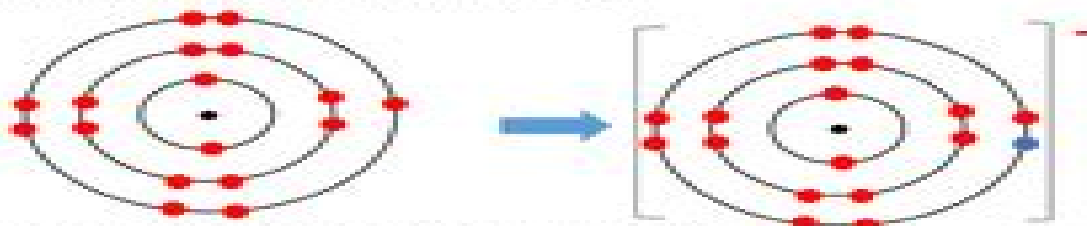


- (g) The diagram shows how a sodium ion is formed from a sodium atom.
 The atomic (proton) number of sodium is 11.



Draw a similar diagram to show how a **chloride ion** is formed from a **chlorine atom**.
 The atomic (proton) number of chlorine is 17.

(3)



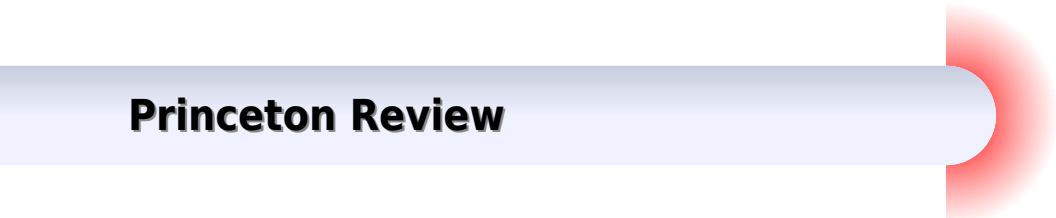
- (h) Explain how the ions are held together in solid sodium chloride.

..... **Positive and negative ions** (1)

..... **Held together by electrostatic forces** (1)

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