

Question

14. Following the redox reaction between silver and copper, write balanced chemical equations for each of the following half-reactions:

Half-reaction	Reaction
1a) Silver ion reduction	$\text{Ag}^+ + e^- \rightarrow \text{Ag}$
1b) Silver metal oxidation	$\text{Ag} \rightarrow \text{Ag}^+ + e^-$
1c) Copper ion reduction	$\text{Cu}^{2+} + 2e^- \rightarrow \text{Cu}$
1d) Copper metal oxidation	$\text{Cu} \rightarrow \text{Cu}^{2+} + 2e^-$
1e) Copper ion reduction	$\text{Cu}^{2+} + 2e^- \rightarrow \text{Cu}$
1f) Copper metal oxidation	$\text{Cu} \rightarrow \text{Cu}^{2+} + 2e^-$
1g) Silver ion reduction	$\text{Ag}^+ + e^- \rightarrow \text{Ag}$
1h) Silver metal oxidation	$\text{Ag} \rightarrow \text{Ag}^+ + e^-$
1i) Copper ion reduction	$\text{Cu}^{2+} + 2e^- \rightarrow \text{Cu}$
1j) Copper metal oxidation	$\text{Cu} \rightarrow \text{Cu}^{2+} + 2e^-$

15. Give three chemical reactions that demonstrate the chemical properties of silver. Write balanced chemical equations for each reaction. (Note: The chemical properties of silver are listed in the table below.)

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G Thomas



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