

Given:  $R(x) = 2x$  and  $G(x) = x^2 + 2$

Find  $\frac{F}{G}(-1)$ .

☒ -2/3

☐ -2

☐ 2/3

NEXT QUESTION

 ASK FOR HELP

# Ignitia Answer Key For Teacher

**Raffaela Di Napoli**



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web hess s law extra ws name answer key 1 find the  $\Delta h$  for the reaction below given the following reactions and subsequent  $\Delta h$  values  
 $2 \text{ CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g}) + 2 \text{ C}_2\text{H}_2(\text{g}) + 5 \text{ O}_2(\text{g}) + 2 \text{ H}_2(\text{g}) + \text{C}_2\text{H}_6(\text{g}) \rightarrow \Delta h = 94.5 \text{ kJ}$   
 $\text{H}_2\text{O}(\text{g}) + \text{H}_2(\text{g}) + \frac{1}{2} \text{ O}_2(\text{g}) \rightarrow \Delta h = 71.2 \text{ kJ}$   
 $\text{C}_2\text{H}_6(\text{g}) + 7 \text{ O}_2(\text{g}) + 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{g}) \rightarrow \Delta h = 283 \text{ kJ}$   
 $\text{C}_2\text{H}_6(\text{g})$

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web chemistry 120 hess s law worksheet calculate  $\Delta h$  for the reaction  $\text{C}_2\text{H}_4(\text{g}) + 2 \text{ C}_2\text{H}_6(\text{g})$  from the following data  
 $\text{C}_2\text{H}_4(\text{g}) + 3 \text{ O}_2(\text{g}) + 2 \text{ CO}_2(\text{g}) + 2 \text{ H}_2\text{O}(\text{l}) \rightarrow \Delta h = 1411 \text{ kJ/mole}$   
 $\text{C}_2\text{H}_6(\text{g}) + 7 \text{ O}_2(\text{g}) + 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{l}) \rightarrow \Delta h = 1560 \text{ kJ/mole}$   
 $\text{H}_2(\text{g}) + \frac{1}{2} \text{ O}_2(\text{g}) + \text{H}_2\text{O}(\text{l}) \rightarrow \Delta h = 285 \text{ kJ/mole}$

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 $\text{C}_2\text{NO}_2(\text{g}) + 1 \text{ O}_2(\text{g}) \rightarrow \text{N}_2\text{O}_5(\text{g})$   
 given the following information

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web libretexts states hess s law states that regardless of the multiple stages or steps of a reaction the total enthalpy change

for the reaction is the sum of all changes  
 Hess's law is written as  $\Delta H_{\text{target}} = \Delta H_1 + \Delta H_2 + \Delta H_3$  or  $\Delta H_{\text{target}} = \sum \Delta H_{\text{known}}$

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web Hess's law states that if a process can be expressed as the sum of two or more steps the enthalpy change for the overall process is the sum of the  $\Delta H$  values for each step to use Hess's law two principles must be understood one if an equation is reversed the sign of the  $\Delta H$  value is also reversed two if an equation is multiplied by a

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web Hess's law states that the total enthalpy change for the reaction will be the sum of all those changes no matter how many different steps or stages in the reaction there are Cohen 2016 the equations for the reactions in the experiment done are as follows  $\text{NaOH} + \text{H}^+ \rightarrow \text{H}_2\text{O}$

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web Hess's law states that if a reaction is the sum of two or more other reactions the  $\Delta H$  for the overall process must be the sum of the  $\Delta H$  values of the constituent reactions enthalpy change  $\Delta H$  is independent of the path that a reaction follows to move from reactants

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web this activity provides a demonstration of Hess's law using three reactions the solubility of  $\text{NaOH}$  in water the solubility of  $\text{NaOH}$  in  $\text{HCl}$  and the reaction of a solution of  $\text{HCl}$  and a solution of  $\text{NaOH}$

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web Hess's law worksheet answers 1 calculate  $\Delta H$  for the reaction  $\text{C}_2\text{H}_4 + \text{H}_2 \rightarrow \text{C}_2\text{H}_6$  from the following data  $\text{C}_2\text{H}_4 + 3\text{O}_2 \rightarrow 2\text{CO}_2 + 2\text{H}_2\text{O}$   $\Delta H = 1411 \text{ kJ}$   $\text{C}_2\text{H}_6 + \frac{7}{2}\text{O}_2 \rightarrow 2\text{CO}_2 + 3\text{H}_2\text{O}$   $\Delta H = 1560 \text{ kJ}$

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web mar 13 2023 be able to use hess s law to calculate reaction enthalpies first law of thermodynamics energy is the ability to do work or transfer heat work is the transfer of energy from one body to another in a sense work is energy in the process of transfer

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web may 8 2015 hess law lab by maya parks partners ben seufert kelsea floyd convert joules to kj in your final answer multiply the mass by the change in temperature and the c given trial 1 use hess s law and the following equations and  $\Delta h$  values to determine the heat of reaction for the reaction c 2 h 4

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web thus for reaction 9 2 in which 1 mole of  $\text{H}_2\text{SO}_4$  and 1 mole of  $\text{NaOH}$  combine to form 2 moles of  $\text{H}_2\text{O}$   $\Delta H = -55.8 \text{ kJ}$  as shown note that the enthalpy change  $\Delta H_{\text{products}} - \Delta H_{\text{reactants}}$  is positive if heat is absorbed that is if  $\Delta H_{\text{products}} > \Delta H_{\text{reactants}}$

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web objective the purpose of the lab is to combine two reactants in the calorimeter and measure the heats of reaction in solution eventually allowing an individual to prove or disprove hess s law introduction thermodynamics is the study of energy exchanged between the system and the surroundings

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web 3 12 hess s law in this experiment you will measure the amount of heat released in these three related exothermic reactions  $\text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$   $\text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O}$   $\text{NaOH} + \text{H}_3\text{PO}_4 \rightarrow \text{Na}_3\text{PO}_4 + \text{H}_2\text{O}$

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