

Chemistry Notes - Grade 10

5.1 Properties and Changes

Physical Change: Water to Ice

Chemical Change: Hydrogen peroxide to Water

Physical Property:

- ↳ Characteristic/description of the substance
- ↳ Based on Qualitative Observations (senses)
- ↳ Also based on Quantitative Observations (calculations and measurements) (e.g. colour, texture, density, smell)

Chemical Property:

- ↳ Describes how a substance behaves in the presence of another substance (e.g. how it reacts with other elements and compounds)
- ↳ Examples are flammability, corrosion, bleaching, rusting etc.

Physical Change:

- ↳ No new substance created
- ↳ Most can be reversed
- ↳ E.g. Change of state, dissolving sugar in water, cutting paper

Chemical Change:

- ↳ Produces a new substance
- ↳ Most *can't* be reversed
- ↳ E.g. Forest fire, boiling an egg, burning wood

Five Clues that a Chemical Change has Occurred:

1. A new colour is observed
2. A new odour is produced
3. Heat or light is produced or absorbed
4. Bubbles of gas are formed
5. A solid material (precipitate) forms in a liquid

5.4 Patterns in the Periodic Table

Where are the metals, non metals, and metalloids located on the periodic table?

- ↳ *Metals:* Left side
- ↳ *Metalloid:* On the staircase
- ↳ *Non Metals:* Right side exception being Hydrogen

Where are the alkali metals found? What are their properties?

- ↳ *Location:* Group 1
- ↳ *Properties:* Highly Reactive, shiny, soft, low melting point, and they have one valence electrons, soft, ductile

Grade Chemistry Notes

Roman Wölfel

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the author's name.

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Revision Notes for Standard Grade Chemistry B. A. Duguid, D. R. Glasby, 1991 *Standard Grade Chemistry Revision Notes* Ellinore Moug, 1992 Higher Grade Chemistry Revision Notes Sandy Herd, Chris Sparling, 1996-01 Revision Notes for Higher Grade Chemistry D. A. Buchanan, 1994 **Standard Grade Chemistry Revision Notes** Sandy Herd, 2000

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