

Name: _____

Date: _____

Student Exploration: Covalent Bonds

Vocabulary: covalent bond, diatomic molecule, Lewis diagram, molecule, noble gases, nonmetal, octet rule, shell, valence, valence electron

Prior Knowledge Questions (Do these BEFORE using the Gizmo.)

1. There are eight markers in a full set, but Flora and Frank each only have seven markers. Flora is missing the red marker, and Frank is missing the blue marker.

What can they do so that each has a full set of markers? **They can share the markers.**

2. Otto and Olivia each have six markers. Otto is missing the purple and green markers, and Olivia is missing the black and brown markers. What can they do so that each has a full set?

They can give the ones they need to each other.

Gizmo Warm-up

Just like the students described above, **nonmetal** atoms can share electrons. As you will see in the *Covalent Bonds* Gizmo, atoms form bonds in this way.

To begin, check that **Fluorine** is selected from the **Select a substance** menu. Click **Play** (▶) to see the electrons orbiting the nucleus of each atom.



1. The outermost electrons in each atom are called **valence electrons**. How many valence electrons does each fluorine atom have? _____7_____
2. Click **Pause** (⏸). Drag an electron from the left atom to the right atom. Click **Play**.
What happens? **They share an electron.**
3. Click **Pause**, drag an electron from the right atom to the left, and then click **Play**.
What happens now? **They share another one.**

Gizmos Covalent Bonds Answers

Shirley Soto

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

Gizmos Covalent Bonds Answers:

Covalent Bond 35 Success Secrets - 35 Most Asked Questions on Covalent Bond - What You Need to Know

Shirley Soto, 2014-10-14 A Blue Ribbon Covalent bond Guide A covalent bond is a biochemical bond that includes the parting of negatron matches amid particles The steady level of alluring and hideous drives amid particles once they share electrons is recognized like covalent joining For numerous particles the parting of electrons permits every one particle to attain the equal of a complete outside shell comparable to a steady microelectronic arrangement There has never been a Covalent bond Guide like this It contains 35 answers much more than you can imagine comprehensive answers and extensive details and references with insights that have never before been offered in print Get the information you need fast This all embracing guide offers a thorough view of key knowledge and detailed insight This Guide introduces what you want to know about Covalent bond A quick look inside of some of the subjects covered Chemical bonding Covalent bond Noncovalent bonding Drug Design Covalent bond History Noncovalent bonding Cation Anion Noncovalent bonding Hydrophobic effect Polar covalent bond Polarity of bonds Noncovalent bonding Boiling Points of Liquids Noncovalent bonding London Dispersion Forces Coordinate Covalent Bond Examples Noncovalent bonding effects Noncovalent bonding H bonding Noncovalent bonding Polar Polar covalent bond Polar molecules Covalent bonds Polarity of covalent bonds Polar covalent bond Polarity of molecules Noncovalent bonding Interaction Noncovalent bonding Van der Waals Forces Covalent bonds Subdivision of covalent bonds Covalent Bond Classification Polar covalent bond Hybrids Noncovalent bonding Electrostatic Interactions Polar covalent bond Nonpolar molecules Covalent Bond Classification Other uses Covalent bonds History Noncovalent bonding Dipole Dipole Noncovalent bonding Protein Folding Structure and much more *Chemical Bonding and the Geometry of Molecules* George E. Ryschkewitsch, 1963 **The Covalent Bond** Henry Sinclair Pickering, 1977

Chemistry of the Covalent Bond Leallyn Burr Clapp, 1957 *A Chemist's Guide to Valence Bond Theory* Sason S. Shaik, Philippe C. Hiberty, 2007-12-04 This reference on current VB theory and applications presents a practical system that can be applied to a variety of chemical problems in a uniform manner After explaining basic VB theory it discusses VB applications to bonding problems aromaticity and antiaromaticity the dioxygen molecule polyradicals excited states organic reactions inorganic organometallic reactions photochemical reactions and catalytic reactions With a guide for performing VB calculations exercises and answers and numerous solved problems this is the premier reference for practitioners and upper level students *Structure and Bonding* Jack Barrett, 2001 This book explains in non mathematical terms where possible the factors that govern covalent bond formation the lengths and strengths of bonds and molecular shapes **Molecules and the Chemical Bond** Henry A. Bent, 2013-09-30 MOLECULES and the Chemical Bond is about understanding Schrdingers equation for chemical systems In his famous Lectures on Physics Richard Feynman quotes Paul Dirac on what it means to understand an equation I understand what an equation means said Dirac if I have a way of figuring out the characteristics of

its solutions without actually solving it That hits the nail on the head Its precisely what Conceptual Valence Bond Theory does for Schrdingers equation A physical understanding of an equation adds Feynman is a completely unmathematical imprecise and inexact thing but absolutely necessary for a physicist It unfolds in MCB in two stages described by Newton as a stage of Analysis a union of observations and inductions and a stage of Synthesis use of inductions accepted as first principles to explain observations The books chief vehicle for creating an intuitive understanding of solutions of Schrdingers equation is the worlds largest and to the authors knowledge virtually only library of line drawings of exclusive orbital models of chemical species electron density profiles By focussing attention on fundamental physical principles and by avoiding use of atomic orbitals and thereby mathematical complexities associated with Schrdingers equation the only source of atomic orbitals the books essays provide a scientifically sound student friendly introduction to modern valence theory Repetition of fundamental ideas here and there is intended to make individual essays understandable and interesting each by itself so that readers may examine them in any order in leisurely walks so to speak in the big garden that is valence theory picking bouquets to their liking

Atoms, Chemical Bonds and Bond Dissociation Energies Sandor Fliszar, 2014-01-15 Chemical bonds their intrinsic energies in ground state molecules and the energies required for their actual cleavage are the subject of this book The theory modelled after a description of valence electrons in isolated atoms explains how intrinsic bond energies depend on the amount of electronic charge carried by the bond forming atoms It also explains how bond dissociation depends on these charges While this theory vividly explains thermochemical stability future research could benefit from a better understanding of bond dissociation if we learn how the environment of a molecule affects its charges we also learn how it modifies bond dissociation in that molecule This essay is aimed at theoretical and physical organic chemists who are looking for new perspectives to old problems

Chemistry of the Covalent Bond Lealyn B. Clapp, 1950 *Acids, Bases, and the Chemistry of the Covalent Bond* Calvin Anthony VanderWerf, 1961 Atomic Structure and Chemical Bonding, a Non-mathematical Introduction Fritz Seel, 1963 *Deciphering the Chemical Code* N. D. Epiotis, 1996 This groundbreaking work the culmination of more than 10 years of research presents a breakthrough theory of chemical bonding across the periodic table Professor Epiotis an internationally known and respected member of the theoretical community challenges the conventional chemical concepts that underlie popular theories of chemical bonding Building on his insight that electron electron repulsion is the single crucial variable that differentiates one chemical system from another the author formulates explains and applies a new approach based on nonorthogonal valence bond methodology that amounts to nothing less than a revolutionary unified theory of chemical bonding across the periodic table This work represents the first post Pauling theory of chemical bonding New theory means new formulae and this work is about new chemical formulae that lead to the self consistent rationalization of existing facts and even more important the design of new chemistry

Answer Book, Chemistry, a Life Science Approach Stuart J. Baum, 1973 Chemical Bonding Clarified Through Quantum Mechanics

George C. Pimentel, Richard D. Spratley, 1969 *Polar Covalence* R Sanderson, 1983-03-28 Polar Covalence provides a detailed account of a successful approach to understanding chemistry from knowledge of atomic structure and the properties that result from this structure This book discusses the nature of multiple bonds Organized into 16 chapters this book begins with an overview of the interrelationships of various basic atomic properties This text then describes chemical bonding which can only occur when the nuclei of both atoms can attract the same electrons Other chapters consider the bond energy of multiple bonds which can be determined by calculating the energy in the usual way as though the bonds were single but of the experimental length This book discusses as well the reduction of the lone pair bond weakening effect through the formation of multiple bonds The final chapter deals with the relative roles of principles and practice in the teaching of inorganic and general chemistry This book is a valuable resource for chemists and students **Chemical Bonding and the Geometry of Molecules** George Eugene Ryschkewitsch, 1965 Covalent Bonding in Crystals, Molecules, and Polymers James C. Phillips, 1969 **Chemical Bonds in Semiconductors and Solids** N. N. Sirota, 1995-12-31

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