

Section 7.2: Mass Relationships in Chemical Reactions

Tutorial 1 Practice, page 323–324

1. (a) Given: $m_{\text{CaCO}_3} = 0.50 \text{ g}$

Required: mass of hydrochloric acid, m_{HCl}

Solution:

Step 1. List the given value, the required value, and the corresponding molar masses.



0.50 g m_{HCl}
100.09 g/mol 36.46 g/mol

Step 2. Convert mass of calcium carbonate to amount of calcium carbonate.

$$n_{\text{CaCO}_3} = 0.50 \text{ g} \times \frac{1 \text{ mol}_{\text{CaCO}_3}}{100.09 \text{ g}}$$

$$n_{\text{CaCO}_3} = 0.004\,996 \text{ mol [two extra digits carried]}$$

Step 3. Convert amount of calcium carbonate to amount of hydrochloric acid.

$$n_{\text{HCl}} = 0.004\,996 \text{ mol}_{\text{CaCO}_3} \times \frac{2 \text{ mol}_{\text{HCl}}}{1 \text{ mol}_{\text{CaCO}_3}}$$

$$n_{\text{HCl}} = 0.009\,992 \text{ mol [two extra digits carried]}$$

Step 4. Convert amount of hydrochloric acid to mass of hydrochloric acid.

$$m_{\text{HCl}} = (0.009\,992 \text{ mol}) \left(\frac{36.46 \text{ g}}{1 \text{ mol}} \right)$$

$$m_{\text{HCl}} = 0.36 \text{ g}$$

Statement: 0.50 g of calcium carbonate will neutralize 0.36 g of hydrochloric acid.

(b) Given: $m_{\text{CaCO}_3} = 0.50 \text{ g}$; $n_{\text{CaCO}_3} = 0.004\,996 \text{ mol}$

Required: mass of calcium chloride, m_{CaCl_2}

Solution:

Step 1. List the given values, the required values, and the corresponding molar masses.



0.004 996 mol n_{CaCl_2}
0.50 g m_{CaCl_2}
100.09 g/mol 110.98 g/mol

Step 2. Convert amount of calcium carbonate to amount of calcium chloride.

$$n_{\text{CaCl}_2} = n_{\text{CaCO}_3}$$

$$n_{\text{CaCl}_2} = 0.004\,996 \text{ mol}$$

Mass Relationships In Chemical Reactions Practice Test

F Rizvi

A decorative graphic consisting of a light blue horizontal bar with a rounded right end, and a red circular shape partially visible behind it.

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