

3.1 - The Periodic Table

3.1.1 - Describe the arrangement of elements in the periodic table in order of increasing atomic number

Elements in the periodic table are arranged in order of **increasing atomic number (Z)**. There is a division between metals and non-metals. Metals are on the left and non-metals are on the right. Metals tend to have a smaller number of electrons in their outer shell

The long metal periods are divided into the **transition metals, lanthanides and actinides**

Hydrogen is difficult to place, however it is placed in group 1 because it displays some of the same characteristics as these elements, although it is a non-metal

Helium is still a noble gas because its outer shell is filled by 2 electrons

1	2											3	4	5	6	7	8									
1 H 1.01		<div>Atomic number Element Atomic mass</div>																								
2 He 4.00																										
3 Li 6.94	4 Be 9.01											5 B 10.81	6 C 12.01	7 N 14.01	8 O 16.00	9 F 18.99	10 Ne 20.18									
11 Na 22.99	12 Mg 24.31											13 Al 26.98	14 Si 28.09	15 P 30.97	16 S 32.06	17 Cl 35.45	18 Ar 39.95									
19 K 39.10	20 Ca 40.08	21 Sc 44.96	22 Ti 47.88	23 V 50.94	24 Cr 52.00	25 Mn 54.94	26 Fe 55.85	27 Co 58.93	28 Ni 58.71	29 Cu 63.55	30 Zn 65.37	31 Ga 69.72	32 Ge 72.64	33 As 74.92	34 Se 78.96	35 Br 79.90	36 Kr 83.80									
37 Rb 85.47	38 Sr 87.62	39 Y 88.91	40 Zr 91.22	41 Nb 92.91	42 Mo 95.94	43 Tc 98.91	44 Ru 101.07	45 Rh 102.91	46 Pd 106.92	47 Ag 107.87	48 Cd 112.41	49 In 114.82	50 Sn 118.71	51 Sb 121.76	52 Te 127.60	53 I 126.90	54 Xe 131.29									
55 Cs 132.91	56 Ba 137.33	57 La 138.91	58 Ce 140.12	59 Pr 140.91	60 Nd 144.24	61 Pm 144.91	62 Sm 150.36	63 Eu 151.96	64 Gd 157.25	65 Tb 158.93	66 Dy 162.50	67 Ho 164.93	68 Er 167.26	69 Tm 168.93	70 Yb 173.05	71 Lu 174.97										
87 Fr (223)	88 Ra (226)	89 Ac (227)											91 Th 232.04	92 Pa 231.04	93 U 238.03	94 Np 237.05	95 Pu 244.06	96 Am 243.06	97 Cm 247.07	98 Bk 247.07	99 Cf 251.08	100 Es 252.08	101 Fm 257.10	102 Md 258.10	103 No 259.10	104 Lr 262.10

Metals

Non-Metals



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Daniel F McAuley



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