# PRACTICE PROBLEMS ON NET IONIC EQUATIONS

Answer Key to Practice Problems on Net Ionic Equations:

Molecular: AgNO<sub>3</sub> (aq) + KCl (aq) → AgCl (s) + KNO<sub>3</sub> (aq)

Total lonic:  $Ag^{-}(aq) + NQ_{0}^{-}(aq) + K^{-}(aq) + Cl^{-}(aq) \rightarrow AgCl(s) + K^{-}(aq) + NQ_{0}^{-}(aq)$ 

Net Ionic: Ag<sup>+</sup>(aq) + Cl<sup>-</sup>(aq) → AgCl(s)

Molecular: Mg(NO<sub>3</sub>)<sub>2</sub> (ag) + Na<sub>2</sub>CO<sub>2</sub> (ag) → MgCO<sub>3</sub> (s) + 2 NaNO<sub>2</sub> (ag)

Total lonic: Mg2 (aq) + 2-NO<sub>4</sub> (aq) + 2-Na\* (aq) + CO<sub>2</sub>\* (aq) → MgCO<sub>3</sub> (s) + 2-Na\* (aq) + 2 NO<sub>4</sub> (aq)

Net lonic:  $Mg^{2-}(aq) + CO_2^{-2}(aq) \rightarrow MgCO_3(s)$ 

Molecular: SrBr₂ (aq) + K₂SO₄ (aq) → SrSO₄ (s) + 2 KBr (aq)

Total lonic:  $Sr^{3}$  (ag) + 2-8r (ag) + 2-K\* (ag) +  $SO_{3}$  (ag)  $\rightarrow$   $SrSO_{4}$  (s) + 2-K\* (ag) + 2-8r (ag)

Net Ionic:  $Sr^{(r)}$  (ag) +  $SO_c^{(r)}$  (ag)  $\rightarrow SrSO_c$  (s)

Molecular: MnCl₂ (aq) + (NH₄)₂CO₃ (aq) → MnCO₃ (s) + 2 NH₄Cl (aq)

Total Ionic: Mn2- (ag) + 2-GI (ag) + 2-NHu- (ag) + CO2- (ag) → MnCO2 (s) + 2-NHu- (ag) + 2-GI (ag)

Net Ionic: Mn2 (ag) + CO₂2 (ag) → MnCO₂ (s)

Molecular: 2 Cr(NO<sub>3</sub>)<sub>3</sub> (aq) + 3 FeSO<sub>4</sub> (aq) → 3 Fe(NO<sub>3</sub>)<sub>3</sub> (aq) + Cr<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub> (aq)

Total lonic;  $2 \text{ Gr}^{\circ}$  (eq) +  $6 \text{ NO}_{\circ}$  (eq) +  $3 \text{ Fe}^{\circ}$  (eq) +  $3 \text{ SO}_{\circ}^{\circ}$  (eq)  $\rightarrow 2 \text{ Gr}^{\circ}$  (eq) +  $6 \text{ NO}_{\circ}$  (eq) +  $3 \text{ Fe}^{\circ}$  (eq) +

3-SQ /- (ag)

Net Ionic: No Reaction

Molecular: K<sub>3</sub>PO<sub>4</sub> (aq) + Al(NO<sub>3</sub>)<sub>3</sub> (aq) → AlPO<sub>4</sub> (s) + 3 KNO<sub>3</sub> (aq)

Total Ionic:  $3 \text{ K}^{\circ}$  (ag) +  $PO_{4}^{\circ}$  (ag) +  $AP^{\circ}$  (ag) +  $3 \text{ NO}_{2}$  (ag)  $\rightarrow$   $AIPO_{4}$  (s) +  $3 \text{ K}^{\circ}$  (ag) +  $3 \text{ NO}_{3}$  (ag)

Net Ionic: AP\* (aq) + PO₂3 (aq) → AIPO₂ (s)

Molecular: Bel<sub>2</sub> (aq) + Cu<sub>2</sub>SO<sub>4</sub> (aq) → BeSO<sub>4</sub> (aq) + 2 Cul (s)

Total Ionic:  $Be^{2r}(aq) + 2I^{-}(aq) + 2Cu^{*}(aq) + SQ_{4}^{2r}(aq) \rightarrow Be^{2r}(aq) + SQ_{4}^{2r}(aq) + 2CuI(s)$ 

Net Ionic:  $2 \text{ Cu}^{\text{pr}}$  (aq) +  $2 \text{ I}^{\text{--}}$  (aq)  $\rightarrow 2 \text{ Cul}$  (s)

Molecular: Ni(NO<sub>3</sub>)<sub>3</sub> (aq) + 3 KBr (aq) -> NiBr<sub>3</sub> (aq) + 3 KNO<sub>3</sub> (aq)

Total lonic; Ni\* (ag) + 3 NO<sub>2</sub> (ag) + 3 K\* (ag) + 3 Br (ag) -> Ni\* (ag) + 3 NO<sub>2</sub> (ag) + 3 K\* (ag) + 3 Br (ag)

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# **A Gutmann**

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