

JHS KAP CHEMISTRY – IONIC BONDING NOTES

CHEMICAL BONDING OVERVIEW

The focus of this unit is on the **IONIC BOND** and naming **IONIC** compounds which are compounds involving atoms bonded together by ionic bonds. We will briefly introduce metallic and covalent bonds as well.

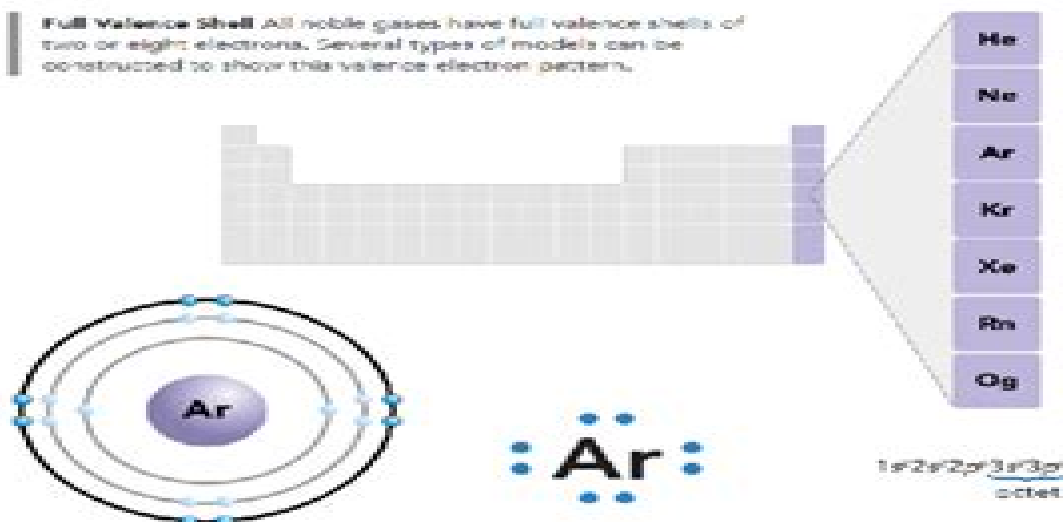
Chemical Bond

A chemical bond is a force of attraction that holds atoms or ions together.

The Octet Rule

- The **octet rule** states that atoms form chemical bonds so that each atom has **eight electrons** (an octet) in its **valence shell** (highest occupied energy level).
- Except for helium which has a total of two electrons, both in the valence shell, the **noble gases** each have **eight electrons** in their **valence shell**. This is an **energetically stable** valence shell electron configuration.
- Atoms of **other elements** achieve the **stable electron configuration** of noble gases by **forming chemical bonds** so that each atom ends up with an **octet** of electrons in their valence shell.
- A few atoms like **hydrogen**, **lithium**, and **beryllium** will become **stable with two electrons** in their valence shell like **helium**.

Full Valence Shell All noble gases have full valence shells of two or eight electrons. Several types of models can be constructed to show this valence electron pattern.



The Bohr model of an argon atom clearly shows that it has eight valence electrons. It also shows the inner electrons.

The electron dot structure shows only the eight valence electrons.

$1s^2 2s^2 2p^6 3s^2 3p^6$
octet

The electron configuration shows that two electrons in the 3s subshell and six in the 3p subshell make up the octet.

DEFINITIONS

If a substance is a **pure solid**, we use the symbol **(s)**. If it is a **pure liquid**, we use the symbol **(l)**. If it is a **pure gas**, we use the symbol **(g)**. When a substance **dissolves in water**, (is soluble in water), we use the symbol **(aq)** to show that it has formed an **aqueous solution**.

Examples:

Solids:	Na_2O	$\text{NaCl}_{(s)}$	$\text{CaCO}_{3(s)}$
Liquids:	$\text{H}_2\text{O}_{(l)}$	$\text{Br}_{2(l)}$	$\text{C}_2\text{H}_5\text{OH}_{(l)}$
Gases:	$\text{CO}_{2(g)}$	$\text{H}_{2(g)}$	$\text{NH}_{3(g)}$
Aqueous solutions:	$\text{LiF}_{(aq)}$	$\text{KNO}_{3(aq)}$	$\text{C}_6\text{H}_{12}\text{O}_{6(aq)}$

See the flow chart on the next page for more definitions.

Ionic Bonding Notes Guide

Konstantinos Papadopoulos

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