
Identify the order of the change that best completes the statement or answers the question.

- ☐ 7. In many compounds, atoms of main-group elements form ions so that the number of electrons in their outermost energy levels of each ion is:
- a. 2; d. 8;
b. 6; e. 10;
- ☐ 8. An atom
- a. is an ion with a negative charge.
b. attracts ions with negative charges.
c. results when an alkaline-earth metal loses one of its two valence-level electrons.
d. has more protons than electrons.
- ☐ 9. When the octet rule is satisfied, the outermost _____ are filled.
- a. all s orbitals b. p and f orbitals
c. s and p orbitals d. d and p orbitals
- ☐ 10. The elements of the _____ group satisfy the octet rule without forming compounds.
- a. halide b. alkali metal
c. noble gas d. alkaline earth metal
- ☐ 11. Which gas atom has full s and p orbitals in its outermost energy level?
- a. It is highly reactive only with alkali metals.
b. It is highly reactive only with halogens.
c. It can be combined with most elements.
d. It has a stable color and is unreactive.
- ☐ 12. In a crystal of an ionic compound, each cation is surrounded by
- a. molecules. b. dipole.
c. positive ions. d. anions.
- ☐ 13. The lithium(I) ion and sodium(I) ion
- a. have the same charge.
b. are polyatomic ions.

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Robert W. Avakian



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