Ionic Compounds and Metals

Section 7.1 Ion Formation

pages 206-209

Section 7.1 Assessment

page 209

 Compare the stability of a lithium atom with that of its ion, Li⁺.

The Li⁺ ion is more stable because it has a complete octet.

Describe two different causes of the force of attraction in a chemical bond.

The two causes are the attraction between the positive nucleus of one atom and the negative electrons of another atom, and the attraction between positive ions and negative ions.

3. Apply Why are all of the elements from group 18 relatively unreactive, whereas those in group 17 are very reactive?

The group 18 elements, known as noble gases, have complete outer energy levels and do not easily form ions. The group 17 elements are highly reactive because they need only to gain a single electron to form an octet.

 Summarize ionic bond formation by correctly pairing these terms: cation, anion, electron gain, electron loss.

(anion, electron gain), (cation, electron loss)

- Apply Write out the electron configuration for each atom. Then predict the change that must occur in each to achieve a noble gas configuration.
 - a. nitrogen

[He]2s³2p³; gain 3 electrons (3 – ion) or lose 5 electrons (5+ ion)

be sufficie

[Ne]3s23p4; gain 2 electrons (2- ion)

c. barium

[Xe]6s2; lose 2 electrons (2+ ion)

d. lithium

[He]2s1; lose 1 electron (1+ ion)

 Model Draw models to represent the formation of the positive calcium ion and the negative bromide ion.

Models should show that the calcium atom loses two electrons, forming Ca²⁺, and that bromine gains one electron, forming Br⁻. The models should also show the addition of energy to form the calcium ion, and the release of energy to form the bromide ion.

Section 7.2 Ionic Bonds and Ionic Compounds

pages 210-217

Practice Problems

page 212

Explain how an ionic compound forms from these elements.

7. sodium and nitrogen

Three Na atoms each lose 1 e⁻, forming 1+ ions. One N atom gains 3 e⁻, forming a 3 – ion. The ions attract, forming Na,N.

3 Ma ions
$$\left(\frac{1+}{\text{Na ion}}\right)$$
 + 1 M ion $\left(\frac{3-}{\text{Nion}}\right)$
= 3(1+) + 1(3-) = 0

The overall charge on one formula unit of Na₃N is zero.

8. lithium and oxygen

Two Li atoms each lose 1 e-, forming 1+ ions. One O atom gains 2 e-, forming a 2- ion. The ions attract, forming Li₂O.

2 Li ions
$$\left(\frac{1+}{\text{Li ions}}\right) + 1 \text{ O-ions}\left(\frac{2-}{\text{O-ions}}\right)$$

= $2(1+) + 1(2-) = 0$

The overall charge on one formula unit of Li₂O is zero.

103

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F Rizvi

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- 1. Understanding the eBook Ionic Compounds And Metals Solutions Manual
 - The Rise of Digital Reading Ionic Compounds And Metals Solutions Manual
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Ionic Compounds And Metals Solutions Manual
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ionic Compounds And Metals Solutions Manual
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Ionic Compounds And Metals Solutions Manual
 - Personalized Recommendations
 - Ionic Compounds And Metals Solutions Manual User Reviews and Ratings
 - Ionic Compounds And Metals Solutions Manual and Bestseller Lists
- 5. Accessing Ionic Compounds And Metals Solutions Manual Free and Paid eBooks
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 - Ionic Compounds And Metals Solutions Manual eBook Subscription Services
 - o Ionic Compounds And Metals Solutions Manual Budget-Friendly Options
- 6. Navigating Ionic Compounds And Metals Solutions Manual eBook Formats

- o ePub, PDF, MOBI, and More
- Ionic Compounds And Metals Solutions Manual Compatibility with Devices
- Ionic Compounds And Metals Solutions Manual Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - o Adjustable Fonts and Text Sizes of Ionic Compounds And Metals Solutions Manual
 - Highlighting and Note-Taking Ionic Compounds And Metals Solutions Manual
 - Interactive Elements Ionic Compounds And Metals Solutions Manual
- 8. Staying Engaged with Ionic Compounds And Metals Solutions Manual
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - $\circ\,$ Following Authors and Publishers Ionic Compounds And Metals Solutions Manual
- 9. Balancing eBooks and Physical Books Ionic Compounds And Metals Solutions Manual
 - Benefits of a Digital Library
 - o Creating a Diverse Reading Collection Ionic Compounds And Metals Solutions Manual
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Ionic Compounds And Metals Solutions Manual
 - Setting Reading Goals Ionic Compounds And Metals Solutions Manual
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ionic Compounds And Metals Solutions Manual
 - Fact-Checking eBook Content of Ionic Compounds And Metals Solutions Manual
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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