

12**INTERPRETING GRAPHICS**Use with Section **6.3****Preparation of Salicylic Acid**

Student #1

mass of flask	37.820 g
flask + $C_2H_4O_2$	39.961 g
volume of $C_2H_4O_2$	5.0 mL
mass of watch glass	22.744 g
watch glass + $C_2H_4O_2$	24.489 g

Student #2

mass of flask	37.979 g
flask + $C_2H_4O_2$	40.010 g
volume of $C_2H_4O_2$	5.0 mL
mass of watch glass	21.688 g
watch glass + $C_2H_4O_2$	24.197 g

Two students prepared aspirin according to the following reaction in which acetic anhydride, $C_4H_6O_3$, reacts with salicylic acid, $C_7H_6O_3$, to form aspirin, $C_9H_8O_4$, and acetic acid, $C_2H_4O_2$.



The procedure involved heating the reaction mixture in a water bath for 15 minutes at 75°C , not to exceed 80°C . The mixture was removed from the water bath, and distilled water was added to decompose any unreacted acetic anhydride. The mixture was then placed in an ice bath for 5 minutes to facilitate the formation of aspirin crystals. The aspirin crystals were collected using filtration. The aspirin crystals were dried and then transferred to a watch glass and massed.

Because their grades were partially based on accuracy, both students used their very best lab technique. Which student got the better grade and why?

1. Determine the molar masses of the following:

a. acetic anhydride, $C_4H_6O_3$, _____

b. salicylic acid, $C_7H_6O_3$, _____

c. aspirin, $C_9H_8O_4$, _____

Interpreting Graphics 12 Perparation Of Salicylic Acid Answer Sheet

Sabine Zange



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Table of Contents Interpreting Graphics 12 Perparation Of Salicylic Acid Answer Sheet

1. Understanding the eBook Interpreting Graphics 12 Perparation Of Salicylic Acid Answer Sheet
 - The Rise of Digital Reading Interpreting Graphics 12 Perparation Of Salicylic Acid Answer Sheet
 - Advantages of eBooks Over Traditional Books
2. Identifying Interpreting Graphics 12 Perparation Of Salicylic Acid Answer Sheet
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Interpreting Graphics 12 Perparation Of Salicylic Acid Answer Sheet
 - User-Friendly Interface
4. Exploring eBook Recommendations from Interpreting Graphics 12 Perparation Of Salicylic Acid Answer Sheet
 - Personalized Recommendations
 - Interpreting Graphics 12 Perparation Of Salicylic Acid Answer Sheet User Reviews and Ratings

- Interpreting Graphics 12 Perparation Of Salicylic Acid Answer Sheet and Bestseller Lists
- 5. Accessing Interpreting Graphics 12 Perparation Of Salicylic Acid Answer Sheet Free and Paid eBooks
 - Interpreting Graphics 12 Perparation Of Salicylic Acid Answer Sheet Public Domain eBooks
 - Interpreting Graphics 12 Perparation Of Salicylic Acid Answer Sheet eBook Subscription Services
 - Interpreting Graphics 12 Perparation Of Salicylic Acid Answer Sheet Budget-Friendly Options
- 6. Navigating Interpreting Graphics 12 Perparation Of Salicylic Acid Answer Sheet eBook Formats
 - ePub, PDF, MOBI, and More
 - Interpreting Graphics 12 Perparation Of Salicylic Acid Answer Sheet Compatibility with Devices
 - Interpreting Graphics 12 Perparation Of Salicylic Acid Answer Sheet Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Interpreting Graphics 12 Perparation Of Salicylic Acid Answer Sheet
 - Highlighting and Note-Taking Interpreting Graphics 12 Perparation Of Salicylic Acid Answer Sheet
 - Interactive Elements Interpreting Graphics 12 Perparation Of Salicylic Acid Answer Sheet
- 8. Staying Engaged with Interpreting Graphics 12 Perparation Of Salicylic Acid Answer Sheet
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Interpreting Graphics 12 Perparation Of Salicylic Acid Answer Sheet
- 9. Balancing eBooks and Physical Books Interpreting Graphics 12 Perparation Of Salicylic Acid Answer Sheet
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Interpreting Graphics 12 Perparation Of Salicylic Acid Answer Sheet
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Interpreting Graphics 12 Perparation Of Salicylic Acid Answer Sheet
 - Setting Reading Goals Interpreting Graphics 12 Perparation Of Salicylic Acid Answer Sheet
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Interpreting Graphics 12 Perparation Of Salicylic Acid Answer Sheet
 - Fact-Checking eBook Content of Interpreting Graphics 12 Perparation Of Salicylic Acid Answer Sheet
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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