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CHAPTER 12**STUDY GUIDE FOR CONTENT MASTERY****Section 12.3 Limiting Reactants**

In your textbook, read about why reactions stop and how to determine the limiting reactant.

Study the diagram showing a chemical reaction and the chemical equation that represents the reaction. Then complete the table. Show your calculations for questions 25–27 in the space below the table.



The molar masses of O_2 , N_2 , and NO are 32.00 g/mol, 28.02 g/mol, and 30.01 g/mol, respectively.

Amount of O_2	Amount of N_2	Amount of NO	Limiting Reactant	Amount and Mass of Excess Reactant
1 molecule	2 molecules	2 molecules	none	none
2 molecules	4 molecules	4 molecules	N_2	2 molecules O_2
3 molecules	6 molecules	6 molecules	3. O_2	3. 4 molecules NO
1.00 mol	2.00 mol	2.00 mol	4. none	4. none
4.00 mol	6.00 mol	4.00 mol	5. NO	5. 2.00 mol O_2
6.00 mol	9.00 mol	6.00 mol	6. N_2	6. 3.00 mol O_2
1.00 mol	6.00 mol	12.00 mol	7. O_2	7. 5.00 mol NO
6.000 mol	9.000 mol	18.000 mol	8. NO	8. 3.000 mol O_2
32.00 g	84.04 g	96.03 g	9. none	9. none
64.00 g	168.08 g	192.06 g	10. O_2	10. 32.00 g NO
96.00 g	252.12 g	288.09 g	11. O_2	11. 48.00 g NO

balanced equation mole ratio = 2 mol NO /1 mol O_2

16.00 g O_2 \times 1 mol O_2 /32.00 g O_2 = 0.500 mol O_2

28.02 g N_2 \times 1 mol N_2 /28.02 g N_2 = 1.000 mol N_2

actual mole ratio = 0.500 mol NO /0.500 mol O_2 = 1.000 mol NO /1.000 mol O_2

Because the actual mole ratio of NO / O_2 is larger than the balanced equation mole ratio of NO / O_2 , there is an excess of N_2 ; O_2 is the limiting reactant.

Mass of NO used = 0.500 mol O_2 \times 2 mol NO /1 mol O_2 = 1.000 mol NO

0.500 mol NO \times 30.01 g NO /1 mol NO = 15.01 g NO

Mass of NO_2 produced = 0.500 mol NO_2 \times 46.01 g NO_2 /1 mol NO_2 = 23.01 g NO_2

Excess N_2 = 28.02 g N_2 - 15.01 g N_2 = 13.01 g N_2

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CHAPTER 13**STUDY GUIDE FOR CONTENT MASTERY****Section 13.4 Percent Yield**

In your textbook, read about the yields of products.

Study the diagram and the example problem.



Example Problem: The following chemical equation represents the production of palladium, a valuable metal in the manufacturing of some semiconductor devices.



In one experiment, the reaction yielded 7.42 g of the metal from a 7.00-g sample of palladium. Determine the percent yield of this reaction. The molar masses of Ga and Ga_2O_3 are 69.72 g/mol and 157.44 g/mol, respectively.

Use the information in the diagram and example problem to evaluate each value or expression below. If the value or expression is correct, write correct. If it is incorrect, write the correct value or expression.

- actual yield unknown 7.42 g Ga_2O_3
- mass of reactant: 7.00 g Ga correct
- number of moles of reactant: 7.00 g Ga \times $\frac{69.72 \text{ g Ga}}{1 \text{ mol Ga}}$ = 1 mol Ga 7.00 g Ga \times 1 mol Ga /69.72 g Ga
- number of moles of product: 0.100 mol Ga \times $\frac{2 \text{ mol Ga}_2\text{O}_3}{1 \text{ mol Ga}}$ = 0.200 mol Ga_2O_3 0.100 mol Ga \times $\frac{2 \text{ mol Ga}_2\text{O}_3}{1 \text{ mol Ga}}$
- theoretical yield: 0.200 mol Ga_2O_3 \times $\frac{157.44 \text{ g Ga}_2\text{O}_3}{1 \text{ mol Ga}_2\text{O}_3}$ = 31.49 g Ga_2O_3 correct
- percent yield: $\frac{7.42 \text{ g Ga}_2\text{O}_3}{31.49 \text{ g Ga}_2\text{O}_3} \times 100\%$ = 23.6% 7.42 g Ga_2O_3 /31.49 g Ga_2O_3 \times 100

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