mudtiply reduction rx. by 3: 3[ClO, (aq) + 8H*(aq) + 8e --- $CI^{-}(aq) + 4H_2O(\ell)$ 3ClO₄"(aq) + 24H*(aq) + 24e" --- $3CI^{-}(aq) + 12H_{2}O(\ell)$ colored block $4Br^{-}(aq) + 12H_{2}\Theta(T) + 3ClO_{4}^{-}(aq) + 24H^{+}(aq)$ + 240 ---4BrO, (aq) + 24H2(aq) + 24g - + 3Cl (aq) + 1214-6900 Burkanica $3ClO_4^-(aq) + 4Br^-(aq) \rightarrow 3Cl^-(aq) + 4BrO_1^-(aq)$ by contiductions: $N_0O_n(g) \rightarrow NO_n^{-1}(aq)$ $N_2O_4(g) \rightarrow 2NO_5^-(mq)$ $N_*O_*(g) + 2H_*O(\ell) \rightarrow 2NO_*^*(aq)$ $N_2O_4(g) + 2H_2O(\ell) \rightarrow 2NO_3^-(aq) + 4H^+(aq)$ $N_2O_\ell(g) + 2H_2O(\ell) + 4OH^*(aq) \rightarrow$ $2NO_3^-(aq) + 4H^+(aq) + 4OH^-(aq)$ $N_2O_4(g) + 2H_2O(\ell) + 4OH^-(aq) \rightarrow$ $2NO_{1}^{-}(sq) + 4H_{2}O(\ell)$ $N_1O_\ell(g) + 2H_2O(\ell) + 4OH^-(aq) \rightarrow$ $2NO_3^-(aq) + 4H_2O(\ell) + 2e^$ reduction: $TeO_s^{2-}(aq) \longrightarrow Te(s)$ $TeO_2^{2-}(sq) \longrightarrow Te(s) + 3H_2O(\ell)$ $TeO_1^{2-}(aq) + 6H^*(aq) + 6OH^*(aq) \rightarrow$ $Te(s) + 3H_2O(\ell) + 6OH^{-}(aq)$ $TeO_s^{1-}(aq) + 6H_sO(\ell) \rightarrow$ $Te(s) + 3H_sO(\ell) + 6OH^-(aq)$ $TeO_s^{2-}(aq) + 3H_sO(\ell) \rightarrow Te(s) + 6OH^-(aq)$ $TeO_3^{2-}(aq) + 3H_2O(\ell) + 4e^- \rightarrow Te(s) + 6OH^-(aq)$ multiply oxidation ex. by 2: $2[N_*O_*(g) + 2H_*O(\ell) + 4OH^*(aq) \rightarrow$ $2NO_{\gamma}^{-}(aq) + 4H_{\gamma}O(\ell) + 2e^{-}$ $2N_zO_z(g) + 4H_zO(\ell) + 8OH^-(aq) \rightarrow$ $4NO_3^{-}(aq) + 8H_2O(\ell) + 4e^{-}$ multiply reduction rs. by I: $TeO_s^{2-}(aq) + 6H_sO(\ell) + 4e^- \rightarrow$ $Te(s) + 3H_2O(\ell) + 6OH^-(aq)$ consultations: $2N_2O_4(g) + 4H_2O(\ell) + 8OH^-(aq) + TeO_1^{1-}(aq) +$ 3/H_(O((f) + _6e**----4NO, (aq) + 8H2O(f) + 4e + Te(s) + 6OH (aq) bullance: $2N_2O_4(g) + TeO_3^{2-}(aq) + 2OH^-(aq) \rightarrow$ $Te(s) + 4NO_{s}^{-}(aq) + H_{s}O(\ell)$ c. excidation: $H_2O_2(aq) \longrightarrow O_2(g)$

 $H_2O_2(aq) \longrightarrow O_2(g) + 2H^+(aq)$

 $H.O.(aq) \rightarrow O.(g) + 2H^*(aq) + 2e^-$

reduction: $Cr_*O_2^{2-}(aq) \rightarrow Cr^{3+}(aq)$ $Cr_*O_*^{2-}(aq) \rightarrow 2Cr^{2+}(aq)$ $Cr_*O_*^{2-}(aq) \rightarrow 2Cr^{2+}(aq) + 7H_*O(\ell)$ $Cr_2O_2^{-2}$ (aq) + 14H*(aq) \rightarrow $2Cr^{3+}(aq) + 7H_2O(\ell)$ Cr₂O₂² (aq) + 14H*(aq) + 6e⁻ → $2Cr^{5+}(\omega a) + 7H_2O(0)$ insultiply oxidation ex. by 3: $3[H_2O_2(aq) \rightarrow O_2(g) + 2H^*(aq) + 2e^-]$ $3H_1O_1(aq) \rightarrow 3O_2(g) + 6H^*(aq) + 6e^$ multiply reduction ex. by 1: Cr₂O₂² (aq) + 14H (aq) + 6q - $2Cr^{5+}(aq) + 7H_2O(\ell)$ consistency. 3H₂O₂(aq) + Cr₂O₂² (aq) + 14H*(aq) + 6e²² -- $3O_2(g) + 6H^*(aq) + 6e^- + 2Cr^{3+}(aq) + 7H_2O(\ell)$ balance: $3H_2O_2(\ell) + Cr_2O_2^{2-}(aq) + 8H^+(aq) \rightarrow$ $3O_{r}(g) + 2Cr^{3+}(sg) + 7H_{r}O(f)$ 14. It is a redox reaction because sulfur is both being

Answers to Section 9.3 Review Questions (Student textbook page 617)

marmber of 0 in S(s).

1. Fluorine is the most electronegative element in the periodic table. Therefore, it does not "lose" electrons in chemical bonds. It has seven valence electrons, and thus forms bonds in which it gains one valence electron, satisfying the octet rule. Because it is in the second row of the periodic table, its outer occupied energy level cannot accommodate more than eight electrons (as bromine, iodine, or chlorine can). Therefore, it does not form bonds in which it gains more than one electron, and its oxidation number is always — I.

exidized from an exidation number of -2 in H₂S(g)

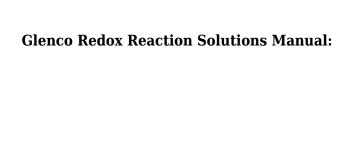
to an oxidation mamber of 0 in S(s) and reduced from

an oxidation number of +4 in SO₂(g) to an oxidation

- 2. a. MacO₂ +4
 - b. +7
 - $c_{-} + 6$
 - dl. $KMmO_n + 7$
- Zero in the diatomic molecule, O₂(g); −1 in peroxides, H₂O₂(aq); −2 in most anions, CO₃^{2−}(aq); +2 in an oxygen difluoride, OF₂(g)
- 4. a. yes
 - the mon
 - C. men
 - dl., yees

Glenco Redox Reaction Solutions Manual

David Kirk



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