Note that the end vowel of the Greek prefix is occasionally omitted from the spelling for clarity in pronunciation.

Hydrates

Prefixes are also used for the naming of hydrates, inorganic salts in which water molecules are a part of the crystalline structure of the solid. The waters of hydration (also called waters of crystallization) are bound to the ions of the salt and, in most cases, can be removed with the application of heat (see Experiment 5). For example, barium chloride is often purchased as burium chloride dihydrate: BaCl₂-2H₂O. The formula of iron(III) chloride Arxehydrate is FeCl₃-6H₂O. The formula for Epsom salt (Figure 5.1) is MgSO₃-7H₂O, magnesium sulfate heptahydrate.

Binary Acids.

A binary acid is an aqueous solution of a compound formed by hydrogen and a more electronegative nonmetal. To name the acid, the prefix hydro- and suffix -ic are added to the root of the nonmetal name.

Formula	Number
HICTORy)	Apaleochilomic acial
Hilber(mg)*	Apublioberomic social
H.Stonet	Apuly-roughthartic sectof

Writing Formulas from Names

In writing formulas for compounds, it is mandatory that the sum of the exidation numbers of the elements equals zero(see Dry Lab 2A, rade 8). The name of the compound must dictate how the formula is to be written.

Example D2B.1 Write the formula for calcium nitride.

Calcium is in group 2A and therefore has an oxidation number of $+2: Ca^{2n}$. The nitrogen atom is always -3 when combined with a metal: N^{3n} . Therefore, for the sum of the oxidation numbers to equal zero, there must be three Ca^{2n} (a total of +6) for every two N^{3n} (a total of -6); the formula is Ca_nN_2 .

DRY LAB Procedure

Procedure Overview: Given the formula of the compound, the proper names for a large number of compounds are to be written. Given the name of the compound, the formulas for a large number of compounds also are to be written.

Your instructor will assign the exercises you are to complete. Answer them on a separate piece of paper. Be sure to indicate the date, your lab section, and your desk number on your Report Sheet. Use the rules that have been described.

- 1. Name the following binary salts of the representative elements.
 - a. Na,P d. CaC, g. Ca,P, j. K,S m. NH,Br p. A)Cl,
 - b. Na₂O e. Cal. h. KCN k. K₂Te n. (NH₂)₂S q. Al₂O₃
 - c. Na₃N f. CaH₂ L KOH L K₂O₃ e. NH₂CN r. AIN
- Name the following salts according to the old -ic, -our system and the Stock system.
 - a. CrS b. Cr₂O₂ c. Crl₂•6H₂O d. CuCl e. Cul₂ f. CuBr₂•4H₂O
- g. HgCl₂ h. Hg₂Cl₂ L HgO
- L HgO J. Fe₂O₃ k. FeS

L. Pel-+6H-O

- n. CoBr₂•6H₂O
 o. SnF₄
- p. SnO₁
 q. Cu₂O

ms. CoO

r. FecOHo

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Carl A. Burtis, David E. Bruns

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