
Question:

Experiment 1: Investigating the effect of temperature on the rate of reaction.

Using a gas syringe, measure the volume of gas produced over time in the reaction between zinc and dilute hydrochloric acid. Repeat the experiment at different temperatures (e.g., 20°C, 30°C, 40°C, 50°C) and record the time taken for a fixed volume of gas to be produced. Plot a graph of time against temperature and determine the rate of reaction at each temperature. Calculate the activation energy of the reaction using the Arrhenius equation.

Answer: 10000 J/mol

Question:

Experiment 2: Investigating the effect of concentration on the rate of reaction.

Using a gas syringe, measure the volume of gas produced over time in the reaction between zinc and dilute hydrochloric acid. Repeat the experiment at different concentrations of hydrochloric acid (e.g., 0.5 M, 1.0 M, 1.5 M, 2.0 M) and record the time taken for a fixed volume of gas to be produced. Plot a graph of time against concentration and determine the rate of reaction at each concentration. Calculate the activation energy of the reaction using the Arrhenius equation.

Answer: 10000 J/mol

Question:

Experiment 3: Investigating the effect of surface area on the rate of reaction.

Using a gas syringe, measure the volume of gas produced over time in the reaction between zinc and dilute hydrochloric acid. Repeat the experiment using different surface areas of zinc (e.g., 1 cm², 2 cm², 4 cm², 8 cm²) and record the time taken for a fixed volume of gas to be produced. Plot a graph of time against surface area and determine the rate of reaction at each surface area. Calculate the activation energy of the reaction using the Arrhenius equation.

Answer: 10000 J/mol

Question:

Experiment 4: Investigating the effect of catalyst on the rate of reaction.

Using a gas syringe, measure the volume of gas produced over time in the reaction between zinc and dilute hydrochloric acid. Repeat the experiment using a catalyst (e.g., copper(II) sulfate) and record the time taken for a fixed volume of gas to be produced. Plot a graph of time against catalyst concentration and determine the rate of reaction at each catalyst concentration. Calculate the activation energy of the reaction using the Arrhenius equation.

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W. Hill, David W. Hall, Marta J.
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caïs torpetius torpes brillant officier et intendant de néron fut sur ordre de ce dernier torturé et décapité à pise pour avoir refusé d abjurer sa foi chrétienne

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