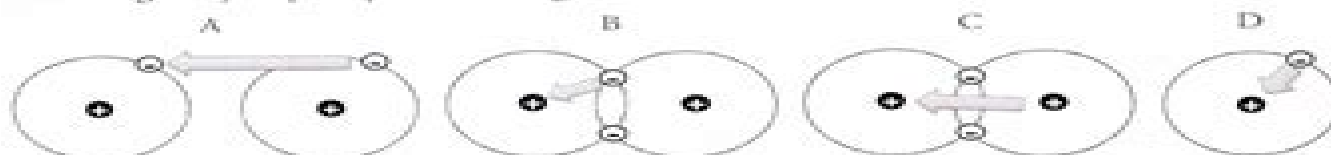


## Read This!

**Electronegativity** is a measure of the ability of an atom's nucleus to attract electrons from a different atom within a covalent bond. A higher electronegativity value correlates to a stronger pull on the electrons in a bond. This value is only theoretical. It cannot be directly measured in the lab.

12. Using the definition stated in the *Read This!* box above, select the best visual representation for electronegativity. Explain your reasoning.



13. Locate the electronegativity values in Model 1.

- a. What is the trend in electronegativity going down a group in Model 1?

*Electronegativity decreases as you go down a group (vertical column).*

- b. Explain the existence of the trend described in part a in terms of atomic structure and Coulombic attraction.

*The ability to attract electrons is based upon attractive force. The greater the distance between the nucleus and valence electrons, the less the attractive force. Less force means a lower electronegativity.*

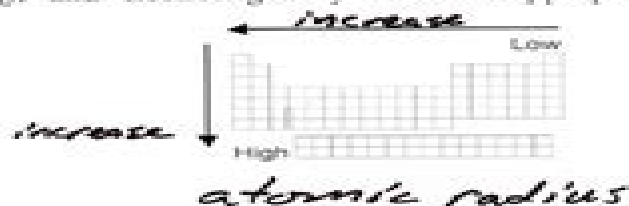
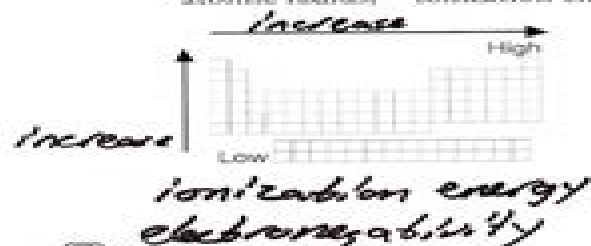
- c. What is the trend in electronegativity going across a period in Model 1?

*Electronegativity increases from left to right across a period (horizontal row)*

- d. Explain the existence of the trend described in part c in terms of atomic structure and Coulombic attraction.

*The greater the number protons in an atom, in a particular period, the greater the attractive force between the nucleus and valence electrons. The greater attractive force means a higher electronegativity.*

14. The two diagrams below can summarize each of the three trends discussed in this activity. Write "atomic radius," "ionization energy," and "electronegativity" under the appropriate diagram.



# High Marks Chemistry Answer Key

**ML Morrison**



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1. Understanding the eBook High Marks Chemistry Answer Key
  - The Rise of Digital Reading High Marks Chemistry Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying High Marks Chemistry Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an High Marks Chemistry Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from High Marks Chemistry Answer Key
  - Personalized Recommendations
  - High Marks Chemistry Answer Key User Reviews and Ratings
  - High Marks Chemistry Answer Key and Bestseller Lists
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  - High Marks Chemistry Answer Key eBook Subscription Services

- High Marks Chemistry Answer Key Budget-Friendly Options
- 6. Navigating High Marks Chemistry Answer Key eBook Formats
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  - High Marks Chemistry Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of High Marks Chemistry Answer Key
  - Highlighting and Note-Taking High Marks Chemistry Answer Key
  - Interactive Elements High Marks Chemistry Answer Key
- 8. Staying Engaged with High Marks Chemistry Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers High Marks Chemistry Answer Key
- 9. Balancing eBooks and Physical Books High Marks Chemistry Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection High Marks Chemistry Answer Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine High Marks Chemistry Answer Key
  - Setting Reading Goals High Marks Chemistry Answer Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of High Marks Chemistry Answer Key
  - Fact-Checking eBook Content of High Marks Chemistry Answer Key
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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