



MERCY JUNIOR COLLEGE
Grade 11
Physical Science
Finals Exam

NAME: _____

YEAR LEVEL & SECTION: _____

TEACHER'S NAME: _____

PROCTOR'S NAME: _____

1. MULTIPLE CHOICE

INSTRUCTION: Encircle the letter of the correct answer. (1 POINT EACH)

- It is a subatomic particle, symbol e^- or β^- , with a negative elementary electric charge.
 - Neutron
 - Electron
 - Proton
 - Atom
- It is a subatomic particle, symbol p or p^+ , with a positive electric charge of $+1e$ elementary charge and mass slightly less than that of a neutron.
 - Neutron
 - Electron
 - Proton
 - Atom
- It is a subatomic particle, symbol n or n^0 , with no net electric charge and a mass slightly larger than that of a proton.
 - Neutron
 - Electron
 - Proton
 - Atom
- Is this an example of ionic bonding?



carbon dioxide

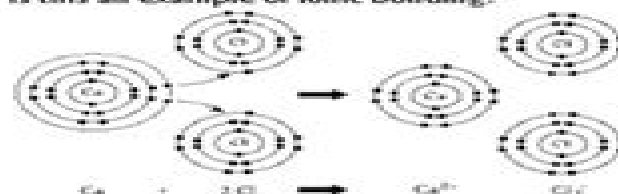
- Yes, it is ionic bonding.
- No, it is of another type of chemical bonding.
- There is no chemical bonding present.
- This is just an orbital model.

- Is this an example of ionic bonding?



lithium fluoride

- Yes, it is ionic bonding.
 - No, it is of another type of chemical bonding.
 - There is no chemical bonding present.
 - This is just an orbital model.
- In atomic physics and quantum chemistry, the electron configuration is the distribution of electrons of an atom or molecule in atomic or molecular orbitals.
 - Electron Configuration
 - Periodic Table of Elements
 - Orbital Models
 - Energy Levels
 - Is this an example of ionic bonding?



- Yes, it is ionic bonding.
- No, it is of another type of chemical bonding.
- There is no chemical bonding present.
- This is just an orbital model.

- It is a chemical rule of thumb that reflects observation that atoms of main-group elements tend to combine in such a way that each atom has eight electrons in its valence shell, giving it the same electronic configuration as a noble gas.
 - Octet Rule
 - Valence Electrons
 - Atomic Orbitals
 - Chemical Bonds
- It is the smallest constituent unit of ordinary matter that has the properties of a chemical element.
 - Neutron
 - Electron
 - Proton
 - Atom
- It is a type of chemical bond that involves the electrostatic attraction between oppositely

Grade12014 Physical Science Final Exam Scope

Puja Mehta



Grade12014 Physical Science Final Exam Scope:

This book delves into Grade12014 Physical Science Final Exam Scope. Grade12014 Physical Science Final Exam Scope is an essential topic that needs to be grasped by everyone, ranging from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Grade12014 Physical Science Final Exam Scope, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:
 - Chapter 1: Introduction to Grade12014 Physical Science Final Exam Scope
 - Chapter 2: Essential Elements of Grade12014 Physical Science Final Exam Scope
 - Chapter 3: Grade12014 Physical Science Final Exam Scope in Everyday Life
 - Chapter 4: Grade12014 Physical Science Final Exam Scope in Specific Contexts
 - Chapter 5: Conclusion
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 3. In chapter 2, the author will delve into the foundational concepts of Grade12014 Physical Science Final Exam Scope. This chapter will elucidate the essential principles that must be understood to grasp Grade12014 Physical Science Final Exam Scope in its entirety.
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 5. In chapter 4, this book will scrutinize the relevance of Grade12014 Physical Science Final Exam Scope in specific contexts. This chapter will explore how Grade12014 Physical Science Final Exam Scope is applied in specialized fields, such as education, business, and technology.
 6. In chapter 5, this book will draw a conclusion about Grade12014 Physical Science Final Exam Scope. The final chapter will summarize the key points that have been discussed throughout the book.
- This book is crafted in an easy-to-understand language and is complemented by engaging illustrations. This book is highly recommended for anyone seeking to gain a comprehensive understanding of Grade12014 Physical Science Final Exam Scope.

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