

- 1) Match the terms in Column A with the descriptions in Column B. Place the letter of your Choice in the space provided. Use your textbook for a reference source if needed. (20)

Column A	COLUMN B
A. Aorta	1. Prevents blood backflow from left ventricle to left atrium
B. Aortic Semilunar Valve	2. Transports blood to left lungs
C. Bicuspid (mitral) Valve	3. Returns venous blood from coronary system to right atrium
D. Chordae Tendineae	4. Major artery leading from left ventricle
E. Inferior Vena Cava	5. Major vein leading into lower right atrium
F. Left Atrium	6. Major vein leading into upper right atrium
G. Left Pulmonary Artery	7. Prevents blood backflow from aorta to left ventricle
H. Left Ventricle	8. Pumps blood to pulmonary artery
I. Myocardium	9. Major artery leading from right ventricle
J. Coronary Arteries	10. Pumps blood to left ventricle
K. Coronary Sinus	11. Major veins leading into left atrium
L. Papillary Muscles	12. Prevents blood backflow from right ventricle to right atrium
M. Pulmonary Artery (trunk)	13. Attach cuspid valve cusps to papillary muscles
N. Pulmonic Semilunar Valve	14. Pumps blood to right ventricle
O. Pulmonary Veins	15. Transport blood from aorta to cardiac muscles
P. Right Atrium	16. Anchor sites for the chordae tendineae
Q. Right Ventricle	17. Pumps blood to aorta
R. Superior Vena Cava	18. Myocardium that separates the two ventricles
S. Tricuspid Valve	19. Prevents blood backflow from pulmonary artery to right ventricle.
T. Ventricular Septum	20. Muscle wall of the heart

- 2) Draw and label a diagram the structure of the human heart showing the proper arrangement or location of EACH of the following structures:

- A) right and left atria;
- B) right and left ventricles;
- C) superior and inferior vena cava;
- D) pulmonary veins;
- E) pulmonary artery;
- F) aorta;
- G) pulmonic and aortic semilunar valves; and,
- H) bicuspid/mitral and tricuspid valves.

Use arrows to indicate the pathway of blood into, through, and out of the heart you have drawn.

(17)

# Grade11 Life Sciences Term4 Practical

**LM Reder**



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web jan 26 2016 the solubility product or ksp is a way of showing how effective a chemical is at dissolving into an aqueous solution the higher the ksp the more the chemical will dissolve into that solvent for this lab calcium hydroxide is created then ionized  $\text{Ca(OH)}_2$  when calcium is put in water reacts as follows  $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$

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web k sp of calcium hydroxide 1 author j m mccormick last update october 13 2013 introduction an equilibrium constant  $K$  is related to  $\Delta G$  for a process through eqn 1 and as long as  $\Delta H$  and  $\Delta S$  are independent of temperature eqn 1 can be combined with eqn 2 to determine  $\Delta H$  and  $\Delta S$  from the temperature dependence of  $K$

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web sep 18 2020 in this experiment we will titrate a saturated calcium hydroxide solution with a standardized hcl solution we will determine the hydroxide ion concentration for the saturated calcium hydroxide

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web your primary objective in this experiment is to test a saturated solution of calcium hydroxide and use your observations and measurements to calculate the ksp of the compound you will do this by titrating the prepared  $\text{Ca(OH)}_2$  solution with a standard hydrochloric acid solution

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web in today's experiment you will determine the solubility product ksp of calcium hydroxide  $\text{Ca(OH)}_2$  by measuring the concentration of  $\text{Ca(OH)}_2$  in a saturated solution calcium hydroxide is a sparingly soluble salt that dissolves according to the following reaction the solubility product expression for this reaction is

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web expert verified chem 182 determining the ksp of calcium hydroxide calcium hydroxide is an ionic solid that is sparingly soluble in water a saturated aqueous solution of  $\text{Ca(OH)}_2$  is represented in equation form as shown below  $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$  the solubility product expression describes in mathematical terms the equilibrium

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web ksp for calcium hydroxide calcium hydroxide is a soft white caustic powder used in making mortar cements calcium salts paints and petrochemicals it is also used in saltwater aquaria to make up kalkwasser limewater solutions for reef tanks and is used as a pH regulating agent notice that calcium hydroxide is divalent and thus has twice

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web use the concentration of calcium hydroxide to calculate the concentration of calcium ions and hydroxide ions in the

solution  $\text{Ca(OH)}_2$  à  $\text{Ca}^{2+}$   $5 \times 10^{-2} \text{ M}$  calcium ions  $5 \times 10^{-2} \text{ M}$   $\text{OH}^-$  hydroxide ions calculate calcium hydroxide K<sub>sp</sub> using the concentrations of calcium ions and hydroxide ions

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web madison bowe elana keller april 15 2014 introduction the objective of this lab was to measure the K<sub>sp</sub> of a slightly soluble basic salt and how solubility changes with temperature K<sub>sp</sub> is another form of equilibrium expression that deals with solubility the reaction that will take place is  $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$

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web question experiment 8 determining K<sub>sp</sub> for calcium hydroxide post lab report sheet data table 2 points trial 1 trial 2 HCl molarity 0.050 M volume of filtered  $\text{Ca(OH)}_2$  solution mL 5.00 5.00 3.71 13.75 equivalence point volume mL OH<sup>-</sup> in saturated  $\text{Ca(OH)}_2$  M Ca<sup>2+</sup> in saturated  $\text{Ca(OH)}_2$  M K<sub>sp</sub> of  $\text{Ca(OH)}_2$  average K<sub>sp</sub> of  $\text{Ca(OH)}_2$

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web the chemical equation used was  $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$  the solubility product was determined to be K<sub>sp</sub>  $\text{Ca(OH)}_2$  from the titration which was made by dripping a concentrated solution of calcium hydroxide into hydrochloric acid the calcium hydroxide was filtered before starting the titration

18.1 solubility product constant K<sub>sp</sub> chemistry libretexts - Feb 01 2023

web jul 12 2023 because the concentration of a pure solid such as  $\text{Ca}_3\text{PO}_4$  is a constant it does not appear explicitly in the equilibrium constant expression the equilibrium constant expression for the dissolution of calcium phosphate is therefore  $\text{Ca}_3\text{PO}_4 \rightleftharpoons 3\text{Ca}^{2+} + 2\text{PO}_4^{3-}$  K<sub>sp</sub>  $\text{Ca}_3\text{PO}_4$

**post lab report 8 determining the K<sub>sp</sub> of calcium hydroxide** - Jul 06 2023

web lab report 8 determining K<sub>sp</sub> of calcium hydroxide the purpose of this was to titrate a saturated  $\text{Ca(OH)}_2$  solution with a standard HCl solution and find the K<sub>sp</sub> or the OH<sup>-</sup> concentration was found and then the concentration was found and used for this experiment a titration of calcium hydroxide with hydrochloric acid

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web question determining the K<sub>sp</sub> of calciumhydroxide  $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$  for this experiment the titration of calciumhydroxide with hydrochloric acid was investigated determining the K<sub>sp</sub> of calcium hydroxide  $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$  for this experiment the titration of calcium

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web therefore it was determined through error calculations and ease of use that the most reliable valid method to use to get the closest ksp value to the theoretical ksp for ca was the determination through ph method 2 introduction calcium hydroxide is a strong base soluble in water

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