

Chemistry Unit 8 Worksheet 4

Samples of Every Kind of Problem

Let us suppose that our "paper" writer is a chemist who has to write out the products of the reaction. Show the given data and find the mass. The mass is underlined (underlined) given.

1. Chlorine gas reacts with a metal to produce metal chloride. MgCl_2 and Mg react. How much chlorine gas is produced per mole of chlorine? MgCl_2 and Mg react. How much chlorine gas?

Equation:	$\text{Mg} + \text{Cl}_2 \rightarrow \text{MgCl}_2$	1	1	1
Mass:	24.3	71	95.3	
Amount:	1.0	1.0	1.0	
Mass:	24.3	71	95.3	

Chlorine gas
95.3 g/mol

2. How much a displacement reaction, magnesium metal reacts with hydrochloric acid to produce magnesium chloride and hydrogen gas. How much magnesium metal reacts with hydrochloric acid to produce a displacement reaction of 1.0 g of magnesium?

Equation:	$\text{Mg} + 2\text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$	1	2	1	1
Mass:	24.3	73	95.3	2	
Amount:	1.0	2.0	1.0	1.0	
Mass:	24.3	73	95.3	2	

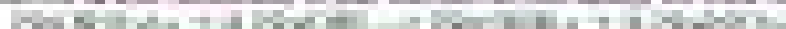
Hydrogen gas
2.0 g/mol

3. Chlorine, Cl_2 , reacts with hydrogen gas to produce hydrogen chloride gas. How much chlorine gas reacts with hydrogen gas to produce a displacement reaction of 1.0 g of hydrogen gas?

Equation:	$\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$	1	1	2	
Mass:	2.0	71	73		
Amount:	1.0	1.0	2.0		
Mass:	2.0	71	73		

Hydrogen gas is 2.0 g/mol. Chlorine gas is 71 g/mol.

4. Chlorine gas reacts with hydrogen gas to produce hydrogen chloride gas. How much chlorine gas reacts with hydrogen gas to produce a displacement reaction of 1.0 g of hydrogen gas?



Chlorine gas is 71 g/mol.

Equation:	$\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$	1	1	2	
Mass:	2.0	71	73		
Amount:	1.0	1.0	2.0		
Mass:	2.0	71	73		

Hydrogen gas is 2.0 g/mol. Chlorine gas is 71 g/mol.

Modeling Chemistry Unit 8 4

R Sanford



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