Classifying and Balancing Equations Multiple Choice

Name Dante PRACTICE TEST

Per

_1. A chemical reaction is a process in which

- a. products change into reactants.
- c. substances can change state
- b. the law of conservation of mass applies d. all of these
- During a chemical reaction,
 - a. new elements are produced
 - b. atoms are rearranged

- c. atoms are destroyed
- d. elements are destroyed
- An equation is balanced by
 - a. changing subscripts

- c. erasing elements as necessary
- b. adding coefficients d. adding elements as necessary
- 4 An atom's ability to undergo chemical reactions is determined by its
 - a. protons b. innermost electrons c. neutrons d. valence electrons
- 5. What are the reactants in the following chemical equation:

zinc and copper

c. zinc and copper (II) sulfate

b. zinc sulfate and copper

- d. only zinc
- 6. What are the products in the above equation?
 - zinc and copper

- c. zinc and copper (II) sulfate
- b. zinc sulfate and copper

d. only zinc

Short Answer

For questions 7-12, classify the reaction according to the type it is. Put that answer in the blank. Then add coefficients to balance the reaction when necessary.

Write a balanced equation for each of the following reactions:

- Magnesium chloride is the product of a reaction between magnesium and chlorine.
- Copper (II) hydroxide and potassium sulfate are produced when potassium hydroxide reacts with copper CID) smilfate.

Unit 6 Practice Test - Key

- L. Dr. 2. % 3. 16
- 4. 45 5. 4

7. Single replacement 1,1,1,1,1 S. Decomposition L.L.L. 9. Double replacement 1,2,1,1-CaCO_{Not} is ppt (see Ref.Table) 10. Single replacement 2,1,1,2

- 11. Combustion 1, 5, 3, 4 12. Combustion 1,3,2,3
- 13. Mg + Cl- -> MgCl- 1.1.1.
- 14. KOH + CuSO₄ -> CutOH)₂ + K₂SO₄ 2,1,1,1

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Robert R. Redfield

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