

## Chemistry Unit 8 Worksheet 4

### Samples of Every Kind of Problem

On a separate sheet of paper, write a complete solution to each of the problems below. Follow the procedure outlined in class. Be sure to circle your final answer.

1. Calculate the number of moles of potassium chlorate,  $\text{KClO}_3$  (s), that must decompose to produce potassium chloride,  $\text{KCl}$  (s), and 1.8 moles of oxygen gas.

|           |   |      |      |                                     |
|-----------|---|------|------|-------------------------------------|
| Equation: | $2 \text{KClO}_3 \rightarrow 2 \text{KCl} + 3 \text{O}_2$ |      |      | 1.2 moles<br>$\text{KClO}_3$ needed |
| Before:   | 1.2   | 0    | 0    |                                     |
| Change:   | -1.2  | +1.2 | +1.8 |                                     |
| After:    | 0   | 1.2  | 1.8  |                                     |

2. In a single displacement reaction, magnesium metal reacts with hydrochloric acid to produce magnesium chloride and hydrogen gas. How many moles of hydrochloric acid are needed to completely react with 2.43 g of magnesium?

|           |   |          |         |       |                                    |
|-----------|---|----------|---------|-------|------------------------------------|
| Equation: | $\text{Mg} + 2 \text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$ |          |         |       | 0.200 moles<br>$\text{HCl}$ needed |
| Before:   | 0.100   | ?        | 0       | 0     |                                    |
| Change:   | -0.100  | -0.200 + | 0.100 + | 0.100 |                                    |
| After:    | 0   | ?        | 0.100   | 0.100 |                                    |

3. Ethane,  $\text{C}_2\text{H}_6$  reacts with oxygen gas to produce carbon dioxide gas and water vapor. What mass of oxygen gas is required to react with 2.20 moles of ethane?

|           |  |       |       |       |
|-----------|--|-------|-------|-------|
| Equation: | $2 \text{C}_2\text{H}_6 + 7 \text{O}_2 \rightarrow 4 \text{CO}_2 + 6 \text{H}_2\text{O}$ |       |       |       |
| Before:   | 2.20   | xs    | 0     | 0     |
| Change:   | -2.20  | -7.70 | +4.40 | +6.60 |
| After:    | 0  | xs    | 4.40  | 6.60  |

$$\text{O}_2 \text{ needed: } 7.70 \text{ mole} \times \frac{32.0 \text{ g}}{1 \text{ mole}} = 246 \text{ g O}_2 \text{ required}$$

4. Determine the mass of sodium nitrate produced when 0.73 g of nickel (II) nitrate reacts with sodium hydroxide according to the following equation:



$$0.73 \text{ g} \times \frac{1 \text{ mole}}{183 \text{ g}} = .0040 \text{ mole}$$

|           |   |         |         |         |
|-----------|---|---------|---------|---------|
| Equation: | $\text{Ni}(\text{NO}_3)_2 + 2 \text{NaOH} \rightarrow \text{Ni}(\text{OH})_2 + 2 \text{NaNO}_3$ |         |         |         |
| Before:   | 0.0040  | XS      | 0       | 0       |
| Change:   | -0.0040   | -0.0080 | +0.0040 | +0.0080 |
| After:    | 0   | XS      | 0.0040  | 0.0080  |

$$\text{NaNO}_3 \text{ produced: } 0.0080 \text{ mole} \times \frac{85.0 \text{ g}}{1 \text{ mole}} = 0.68 \text{ g}$$

# Modeling Chemistry Unit 4 4 Answers

**Norman Herr, James Cunningham**



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