

Chemistry – Unit 5 Worksheet 2

1. An old (pre-1987) penny is nearly pure copper. If such a penny has a mass of 3.3 g, how many moles of copper atoms would be in one penny?

$$3.3 \text{ g} \times \frac{1 \text{ mole}}{63.5 \text{ g}} = 5.2 \times 10^{-2} \text{ mol or } 0.052 \text{ mol}$$

2. Four nails have a total mass of 4.42 grams. How many moles of iron atoms do they contain?

$$4.42 \text{ g} \times \frac{1 \text{ mole}}{55.8 \text{ g}} = 7.92 \times 10^{-2} \text{ mol or } 0.0792 \text{ mol}$$

3. A raindrop has a mass of 0.050 g. How many moles of water does a raindrop contain?

$$0.050 \text{ g} \times \frac{1 \text{ mole}}{18.0 \text{ g}} = 2.78 \times 10^{-3} \text{ mol} \rightarrow 2.8 \times 10^{-3} \text{ mol}$$

4. What mass of water would you need to have 15.0 moles of H₂O?

$$15.0 \text{ moles} \times \frac{18.0 \text{ g}}{1 \text{ mole}} = 2.70 \times 10^2 \text{ g}$$

5. One box of Morton's Salt contains 737 grams. How many moles of sodium chloride is this?

$$737 \text{ g} \times \frac{1 \text{ mole}}{58.5 \text{ g}} = 12.6 \text{ mol}$$

Na	23.0
Cl	35.5
	<hr/> 58.5 g

6. A chocolate chip cookie recipe calls for 0.050 moles of baking soda (sodium bicarbonate). How many grams should the chef mass out?

$$0.050 \text{ moles} \times \frac{84.0 \text{ g}}{1 \text{ mole}} = 4.2 \text{ g}$$

Na	23.0
H	1.0
C	12.0
O	48.0
	<hr/> 84.0

7. Rust is iron(III) oxide. The owner of a 1959 Cadillac convertible wants to restore it by removing the rust with oxalic acid, but he needs to know how many moles of rust will be involved in the reaction. How many moles of iron(III) oxide are contained in 2.50 kg of rust?

$$2.50 \text{ kg} \times \frac{1000 \text{ g}}{1 \text{ kg}} \times \frac{1 \text{ mole}}{159.6 \text{ g}} = 15.7 \text{ mol Fe}_2\text{O}_3$$

Fe	111.6
O	48.0
	<hr/> 159.6

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R Sanford



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Table of Contents Modeling Chemistry U5 Quiz 1 V2 0

1. Understanding the eBook Modeling Chemistry U5 Quiz 1 V2 0
 - The Rise of Digital Reading Modeling Chemistry U5 Quiz 1 V2 0
 - Advantages of eBooks Over Traditional Books
2. Identifying Modeling Chemistry U5 Quiz 1 V2 0
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Modeling Chemistry U5 Quiz 1 V2 0
 - User-Friendly Interface
4. Exploring eBook Recommendations from Modeling Chemistry U5 Quiz 1 V2 0
 - Personalized Recommendations
 - Modeling Chemistry U5 Quiz 1 V2 0 User Reviews and Ratings
 - Modeling Chemistry U5 Quiz 1 V2 0 and Bestseller Lists
5. Accessing Modeling Chemistry U5 Quiz 1 V2 0 Free and Paid eBooks

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- Modeling Chemistry U5 Quiz 1 V2 0 Budget-Friendly Options
- 6. Navigating Modeling Chemistry U5 Quiz 1 V2 0 eBook Formats
 - ePub, PDF, MOBI, and More
 - Modeling Chemistry U5 Quiz 1 V2 0 Compatibility with Devices
 - Modeling Chemistry U5 Quiz 1 V2 0 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Modeling Chemistry U5 Quiz 1 V2 0
 - Highlighting and Note-Taking Modeling Chemistry U5 Quiz 1 V2 0
 - Interactive Elements Modeling Chemistry U5 Quiz 1 V2 0
- 8. Staying Engaged with Modeling Chemistry U5 Quiz 1 V2 0
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Modeling Chemistry U5 Quiz 1 V2 0
- 9. Balancing eBooks and Physical Books Modeling Chemistry U5 Quiz 1 V2 0
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Modeling Chemistry U5 Quiz 1 V2 0
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Modeling Chemistry U5 Quiz 1 V2 0
 - Setting Reading Goals Modeling Chemistry U5 Quiz 1 V2 0
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Modeling Chemistry U5 Quiz 1 V2 0
 - Fact-Checking eBook Content of Modeling Chemistry U5 Quiz 1 V2 0
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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