

Unit 2 Worksheet 1: Molar relationships

For each of the problems below:

- Write a full balanced chemical equation.
- Calculate, where it applies, molar masses and select the most useful molar ratio.
- How much of a mass balanced experimental deviation is allowed different sort of compounds?

1. Hydrogen sulfide gas, which smells like rotten eggs, burns in air to give gaseous sulfur dioxide and water. How many moles of oxygen gas would be needed to completely burn 3 moles of hydrogen sulfide?

Equation: $3\text{H}_2\text{S}_{(g)} + \dots \text{O}_{2(g)} \rightarrow \dots \text{SO}_{2(g)} + \dots \text{H}_2\text{O}_{(g)}$

Molar: $3 \quad \quad \quad 3 \quad \quad \quad 3 \quad \quad \quad 3$

Change: $3 \text{ mol} \quad \quad \quad 3 \quad \quad \quad 3 \quad \quad \quad 3$

Answer: $3\text{H}_2\text{S}_{(g)} + 3\text{O}_{2(g)} \rightarrow 3\text{SO}_{2(g)} + 3\text{H}_2\text{O}_{(g)}$

2. Propane, C_3H_8 , burns in air to form carbon dioxide and water. What volume of carbon dioxide can be made from 20 g of propane at STP?

Equation: $\text{C}_3\text{H}_{8(g)} + \dots \text{O}_{2(g)} \rightarrow \dots \text{CO}_{2(g)} + \dots \text{H}_2\text{O}_{(g)}$

Molar: $1 \quad \quad \quad 3 \quad \quad \quad 1 \quad \quad \quad 4$

Change: $20 \text{ g} \quad 2.27 \quad \quad \quad 6.81 \quad \quad \quad 11.0$

Answer: $2\text{C}_3\text{H}_{8(g)} + 10\text{O}_{2(g)} \rightarrow 6\text{CO}_{2(g)} + 8\text{H}_2\text{O}_{(g)}$

3. Aluminum, Al , the lightweight metal, is reacting hydrogen and nitrogen to form an ionic nitrogenous metal compound. How many moles of a nitrogen gas can be made from 10 g of metallic hydrogen gas?

Equation: $\text{H}_{2(g)} + \dots \text{N}_{2(g)} \rightarrow \dots \text{AlN}_{3(s)}$

Molar: $1 \quad \quad \quad 1 \quad \quad \quad 1$

Change: $10 \text{ mol} \quad 0.47 \quad \quad \quad 0.47 \quad \quad \quad 0.47$

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M Mosston



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