

# GRADE 11 MATH (mcr3u)

## FINAL EXAM

**Instructions:** There is a 2 hour minimum, 3 hour maximum time limit for writing this exam.

**Note:** Number of Pages (including cover): 12

Total Marks: 123

**Give exact answers where possible. Round approximate answers to 2 decimal places unless otherwise indicated.**

Please check over your solutions!

Do your best! Do the questions in any order. Don't get stuck on one. Leave it and come back to it later.

Topic of Study	Points
Multiple Choice	16
Unit 1: Functions	24
Unit 2: Rational Expressions	8
Unit 3: Transformations	15
Unit 4: Exponential Functions	15
Unit 5/6: Trigonometry	29
Unit 7: Discrete Functions	16
<b>Total</b>	<b>123</b>

KTAC breakdown	Knowledge	Application	Thinking	Communication
% of exam	36%	31%	24%	12%

# Mathematics Final Exam Grade1p1

**Emilie Sanchez**



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web a reaction and its experimental rate law are represented below  $2 \text{ ICl}_3(g) \rightarrow 2 \text{ I}_2(g) + 2 \text{ HCl}(g)$  rate  $k [\text{ICl}_3]^2$  a chemist proposes a mechanism for the reaction that is consistent with the rate law the mechanism has two elementary steps and the first step is slow compared to the second

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consider the following mechanism  
 $\text{O}_3 \xrightarrow{\text{fast}} \text{O}_2 + \text{O}$   
 $\text{O} + \text{O}_2 \xrightarrow{\text{slow}} \text{O}_3$   
 a) write the overall balanced chemical equation  
 $2 \text{O}_3 \rightarrow 3 \text{O}_2$   
 b) identify any intermediates within the mechanism  
 $\text{O}$

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 $\text{Mg} + 2 \text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$   
 the reaction between  $\text{Mg}$  and  $\text{HCl}$  is represented by the equation above in a kinetics experiment  
 a) 0.080 g sample of  $\text{Mg}$  ribbon is added to an erlenmeyer flask containing 50 ml of 3.0 M  $\text{HCl}$  at 25 °C

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