Question	Acceptable Answers	Reject	Mark
22(b)(iii)	ΔH <sub>reaction</sub> = +2.3(10) x2 = +23.1 (kJ mol <sup>-1</sup> ) 0.2(00) = +23 (kJ mol <sup>-1</sup> ) to 2 sf  First mark - correct computation of ΔH <sub>reaction</sub> :  2 x [answer to (b)(i) in kJ ÷ answer to (b)(ii) in mol]  Second mark - stand alone, for correct rounding:  A final answer to two sf  Third mark - stand alone, for giving a + sign for endothermic reaction: + sign in front of final answer  NOTE: + 12 (kJ mol <sup>-1</sup> ) scores (2) (i.e. the 2nd and 3rd marks)	Incorrect units given by the	3

Question Number	Acceptable Answers	Reject	Mark
22(c)(i)	(Average amount of) energy/enthalpy required to break one mole of (covalent) bonds  ALLOW Energy change/enthalpy change to break one mole of (covalent) bonds  (1)	Energy/enthalpy released OR 'Bonds formed/made' OR 1 mol of compound for 1st mark	2
	(in the) gas / gaseous (state) (1)		

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**Victor M. Corman** 

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