

Grade 12 Chemistry Exam Review

Below is a list of major topics covered on the exam for this course. If a topic is not listed, it will NOT be on the exam.

Your studying should begin with your class notes, assigned problems, textbook problems and extra review questions suggested below. This list is a guide and some specific questions on the exam may not appear here.

Good luck ☺

Unit #1 – Energy Changes Text Reference: Chapter 5

- ☐ Energy changes in reactions (system, surroundings, endothermic vs. exothermic)
- ☐ Specific Heat Capacity
- ☐ Calorimetry, $q = mc\Delta T$
- ☐ Molar enthalpies
- ☐ Representing Enthalpy Changes (Thermochemical Equations, Potential Energy Diagrams, etc.)
- ☐ Hess' Law
- ☐ Standard Enthalpies of Formation

Sample Exam Questions

1. Given the following information:

$$2\text{CH}_3\text{OH}(\text{l}) + 3\text{O}_2(\text{g}) \rightarrow 2\text{CO}_2(\text{g}) + 4\text{H}_2\text{O}(\text{l}) \quad \Delta H^\circ = -727 \text{ kJ}\cdot\text{mol}^{-1}$$

$$\Delta H_f^\circ \text{CO}_2(\text{g}) = -394 \text{ kJ}\cdot\text{mol}^{-1}, \quad \Delta H_f^\circ \text{H}_2\text{O}(\text{l}) = -286 \text{ kJ}\cdot\text{mol}^{-1}$$

What is the standard heat formation, ΔH_f° (kJ·mol⁻¹) of CH₃OH(l)?

A. -602 B. -1329 C. 1329 D. 602
2. When 789 g of liquid ammonia is cooled from 82.7 °C to 25.0 °C it loses 214 kJ of heat. What is the specific heat capacity of liquid ammonia?

A. 0.213 B. 4.19 C. 0.0047 D. 2.06 E. 4.70
3. Identify the correct expressions that can be used to determine the change in thermal energy for the following reaction:

$$\text{C}_4\text{H}_{10}\text{g} + \frac{13}{2}\text{O}_{2\text{g}} \rightarrow 4\text{CO}_{2\text{g}} + 5\text{H}_2\text{O}_{\text{g}} + \text{heat?}$$

A. $\Delta H^\circ = [(4 \text{ mol})(\Delta H^\circ, \text{CO}_{2\text{g}}) + (5 \text{ mol})(\Delta H^\circ, \text{H}_2\text{O}_{\text{g}})] + [(1 \text{ mol})(\Delta H^\circ, \text{C}_4\text{H}_{10\text{g}}) + (\frac{13}{2} \text{ mol})(\Delta H^\circ, \text{O}_{2\text{g}})]$ B. $\Delta H^\circ = [(1 \text{ mol})(\Delta H^\circ, \text{C}_4\text{H}_{10\text{g}}) + (\frac{13}{2} \text{ mol})(\Delta H^\circ, \text{O}_{2\text{g}})] - [(4 \text{ mol})(\Delta H^\circ, \text{CO}_{2\text{g}}) + (5 \text{ mol})(\Delta H^\circ, \text{H}_2\text{O}_{\text{g}})]$

C. $\Delta H^\circ = [(1 \text{ mol})(\Delta H^\circ, \text{C}_4\text{H}_{10\text{g}}) + (\frac{13}{2} \text{ mol})(\Delta H^\circ, \text{O}_{2\text{g}})] - [(4 \text{ mol})(\Delta H^\circ, \text{CO}_{2\text{g}}) + (5 \text{ mol})(\Delta H^\circ, \text{H}_2\text{O}_{\text{g}})]$

D. $\Delta H^\circ = [(4 \text{ mol})(\Delta H^\circ, \text{CO}_{2\text{g}}) + (5 \text{ mol})(\Delta H^\circ, \text{H}_2\text{O}_{\text{g}})] - [(1 \text{ mol})(\Delta H^\circ, \text{C}_4\text{H}_{10\text{g}}) + (\frac{13}{2} \text{ mol})(\Delta H^\circ, \text{O}_{2\text{g}})]$

E. $\Delta H^\circ = [(1 \text{ mol})(\Delta H^\circ, \text{CO}_{2\text{g}}) + (\frac{4}{5} \text{ mol})(\Delta H^\circ, \text{H}_2\text{O}_{\text{g}})] - [(1 \text{ mol})(\Delta H^\circ, \text{C}_4\text{H}_{10\text{g}}) + (\frac{13}{2} \text{ mol})(\Delta H^\circ, \text{O}_{2\text{g}})]$
4. For which of the following statements is correct for the reaction

$$2\text{H}_2\text{O}_2(\text{g}) \rightarrow \text{H}_2(\text{g}) + 2\text{O}_2(\text{g}) + 66.4 \text{ kJ}$$

A. exothermic, potential energy greater in reactants
 B. exothermic, potential energy greater in products
 C. endothermic, potential energy greater in products
 D. endothermic, potential energy less in reactants
 E. endothermic, potential energy less in products

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Victor M. Corman



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Table of Contents Grade 12 Chemistry Exam Tips Study Guide

1. Understanding the eBook Grade 12 Chemistry Exam Tips Study Guide
 - The Rise of Digital Reading Grade 12 Chemistry Exam Tips Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Grade 12 Chemistry Exam Tips Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Grade 12 Chemistry Exam Tips Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Grade 12 Chemistry Exam Tips Study Guide
 - Personalized Recommendations
 - Grade 12 Chemistry Exam Tips Study Guide User Reviews and Ratings
 - Grade 12 Chemistry Exam Tips Study Guide and Bestseller Lists
5. Accessing Grade 12 Chemistry Exam Tips Study Guide Free and Paid eBooks

- Grade 12 Chemistry Exam Tips Study Guide Public Domain eBooks
 - Grade 12 Chemistry Exam Tips Study Guide eBook Subscription Services
 - Grade 12 Chemistry Exam Tips Study Guide Budget-Friendly Options
6. Navigating Grade 12 Chemistry Exam Tips Study Guide eBook Formats
- ePub, PDF, MOBI, and More
 - Grade 12 Chemistry Exam Tips Study Guide Compatibility with Devices
 - Grade 12 Chemistry Exam Tips Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Grade 12 Chemistry Exam Tips Study Guide
 - Highlighting and Note-Taking Grade 12 Chemistry Exam Tips Study Guide
 - Interactive Elements Grade 12 Chemistry Exam Tips Study Guide
8. Staying Engaged with Grade 12 Chemistry Exam Tips Study Guide
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Grade 12 Chemistry Exam Tips Study Guide
9. Balancing eBooks and Physical Books Grade 12 Chemistry Exam Tips Study Guide
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Grade 12 Chemistry Exam Tips Study Guide
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Grade 12 Chemistry Exam Tips Study Guide
- Setting Reading Goals Grade 12 Chemistry Exam Tips Study Guide
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Grade 12 Chemistry Exam Tips Study Guide
- Fact-Checking eBook Content of Grade 12 Chemistry Exam Tips Study Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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