

**Chapter 1: Stoichiometry**

**Q4.**

**Sol:**  $\text{CaCO}_3 \longrightarrow \text{CaO} + \text{CO}_2$

Since 1 mole of  $\text{CaCO}_3$  produces 1 mole of  $\text{CaO}$ , So

1 mole of  $\text{CaO}$  will produce 1 mole of  $\text{CaCO}_3$ .

$\therefore$  1 mole of  $\text{CaO}$   $\longrightarrow$  1 mole of  $\text{CaCO}_3$ .

**Sol:**  $\text{BaCl}_2 \longrightarrow \text{BaCO}_3$

**Given:**

5 moles of  $\text{BaCl}_2$  contains 5 moles of  $\text{Ba}$  & 10 moles of  $\text{Cl}$  & each form 5 moles of  $\text{BaCO}_3$ ,  $\therefore$  each 1 mole of  $\text{BaCO}_3$  contains 1 mole of  $\text{Ba}$ .

**Sol:**  $2 \text{Na} \longrightarrow \text{Na}_2\text{O}$

5 moles  $\text{O}_2$  (given)

$\therefore$  2 moles produces 1 mole of  $\text{Na}_2\text{O}$

$\therefore 5 \longrightarrow \frac{5}{2} \times 2 = 5$

$\therefore 5 \longrightarrow 5 \times \frac{1}{2} = 2.5$  moles of  $\text{Na}_2\text{O}$ .

**Q1.**

**Ans:**  $\text{Ca} + \frac{1}{2} \text{O}_2 \longrightarrow \text{CaO}$  (balanced chemical equation)

Wt. of  $\text{Ca} = 40 \text{ g/mole}$

So, moles of  $\text{Ca} = \frac{80}{40} = 2 \text{ moles}$

From above balanced chemical equation, 2 moles of  $\text{Ca}$  react with  $\frac{1}{2} \times 2$  moles of  $\text{O}_2$ .

Wt. of  $\text{O}_2$  with which 80g of  $\text{Ca}$  react  $= \frac{1}{2} \times 32 = 16 \text{ g}$  **Ans.**

**Q2.**

**Sol:**  $\text{NaCl} \longrightarrow \text{NaClO}_3$

58 gm.

If we have not balanced the chemical equation, then we have to apply PMOL rule or g-equivalent rule.

By PMOL rule Cl atom

From  $\text{NaCl}$  atom to  $\text{NaClO}_3$  atom  $\therefore$  1 mol. of Cl atom to 1 mol. of Cl atom

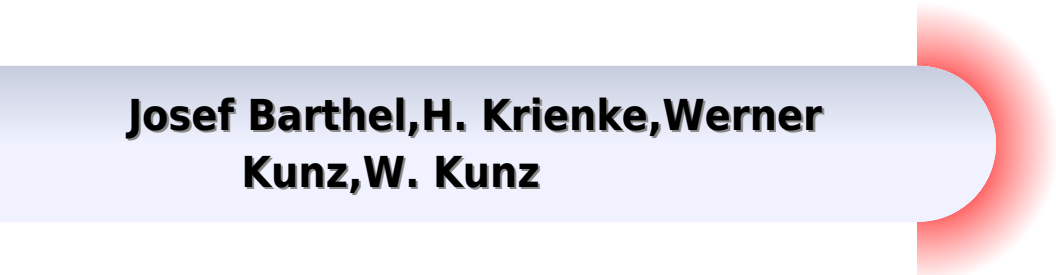
So, 1 mole of  $\text{NaCl}$   $\rightarrow$  1 mole of  $\text{NaClO}_3$   $\therefore$  2 moles of  $\text{NaCl}$   $\rightarrow$  2 moles of  $\text{NaClO}_3$

$\frac{58}{58.5} = 1$  mole of  $\text{NaCl}$

moles of  $\text{Cl}_2 = \frac{58}{117}$

# Modern Approach To Chemical Calculations Solutions

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