

## Chemistry Unit 8 Worksheet 4

### Examples of Every Kind of Problem

This is a separate sheet of paper, so this is a complete solution to each of the problems. Notice: I'll state the problems without an arrow. We have to write your final answer.

12. Calculate the number of moles of potassium chlorate,  $\text{KClO}_3$ , that must decompose to produce potassium chloride  $\text{KCl}$  and a total of 1.5 moles of oxygen gas.



Subst:	1.0	0	0
Change:	-1.0	+1.0	+1.5
Final:	0	1.0	1.5

**1.0 moles  $\text{KClO}_3$  needed**

13. As a single displacement reaction, magnesium metal reacts with hydrochloric acid to produce magnesium chloride and hydrogen gas. How many moles of hydrochloric acid are needed to completely react with 5.1 g of magnesium?



Subst:	0.200	0	0	0
Change:	-0.200	-0.400	+0.200	+0.200
Final:	0	0	0.200	0.200

**0.400 moles HCl needed**

14. Ethane,  $\text{C}_2\text{H}_6$ , reacts with oxygen gas to produce carbon dioxide gas and water vapor. What number of moles of oxygen are required to react with 0.50 moles of ethane?



Subst:	0.50	0.0	0	0
Change:	-0.50	-0.75	+1.00	+1.50
Final:	0	0.25	1.00	1.50

**0.75 moles  $\text{O}_2$  needed** (0.50 moles  $\text{C}_2\text{H}_6$  x 1.50 = 0.75  $\text{O}_2$  needed)

15. Sodium metal that is used in various industries produced 1.00 g of sodium. How many moles of sodium hydroxide are needed for the following reactions:



Subst:	0.0430	0.0215	0	0
Change:	-0.0430	-0.0215	+0.0215	+0.0430
Final:	0	0.0215	0.0215	0.0430

**0.0430 moles of sodium needed** (0.0430 x 1 = 0.0430)

# Modeling Chemistry Unit 8 Test Answers

**L Darling-Hammond**



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