# Chemistry Unit 8 Worksheet 4 Samples of Every Kind of Problem

On a separate sheet of paper, write a complete solution to each of the problems below. Follow the procedure outlined in class. Be sure to circle your final answer.

 Calculate the number of moles of potassium chlorate, KClO<sub>3</sub> (s), that must decompose to produce potassium chloride, KCl (s), and 1.8 moles of oxygen gas.

| Equation: | 2 KClO <sub>3</sub> + | ≥ 2 KCl + | 3 02 |
|-----------|-----------------------|-----------|------|
| Before:   | 1.2                   |           | 0    |
| Change:   | -1.2                  | +1.2      | +1.8 |
| Afters    | .0                    | 1.2       | 1.8  |

1.2 moles KCRO, needed

2. In a single displacement reaction, magnesium metal reacts with hydrochloric acid to produce magnesium chloride and hydrogen gas. How many moles of hydrochloric acid are needed to completely react with 2.43 g of magnesium?

| Equations | Non-   | 2 HCl >  | MgCl <sub>2</sub> + | 882   |
|-----------|--------|----------|---------------------|-------|
| Before:   | 0.100  | 7        |                     | 0     |
| Change:   | -0.100 | -0.200 + | 0.100 +             | 0.100 |
| Afters    |        | 2        | 0.100               | 0.100 |

0.200 protes HCI needed

0.73 g x 1 mole = .0040 mole 183 gr

 Ethane, C<sub>2</sub>H<sub>6</sub> reacts with oxygen gas to produce carbon dioxide gas and water vapor. What mass of oxygen gas is required to react with 2.20 moles of ethane?

| Equation: | 2 C2H6 | + 7 O <sub>2</sub> | → 4 CO₂ | + 6 H <sub>2</sub> O |
|-----------|--------|--------------------|---------|----------------------|
| Before:   | 2.20   | 5.5                | 0       | 0.0                  |
| Change:   | -2.20  | ~ 7.70             | + 4,40  | + 6,60               |
| After:    | 0      | 3.5                | 4.40    | 6.60                 |

O2 needed: 7,70 mole x 32.0 g = 246 g O2 required I mode

 Determine the mass of sodium nitrate produced when 0.73 g of nickel (II) nitrate reacts with sodium hydroxide according to the following equation: Ni(NO3)2 + 2 NaOH --> Ni(OH)2 + 2 NaNO3

> Ni(NO<sub>3</sub>)<sub>2</sub> + 2 NaOH -> Ni(OH)<sub>2</sub> + 2 NaNO<sub>3</sub> Equation: Before: 0.0040 No. 26 -0.0040- O. GOOSIO **-0.0040** +0.06900

N. S. 0.0080 Afflors 0.0640 NaNO<sub>2</sub> produced: 0.0080 mole x 85.0 g = 0.68 g 1 mode

Change:

# **Modeling Chemistry Unit 7 1 Answers**

**PT Brinkman** 

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# **Table of Contents Modeling Chemistry Unit 7 1 Answers**

- 1. Understanding the eBook Modeling Chemistry Unit 7 1 Answers
  - The Rise of Digital Reading Modeling Chemistry Unit 7 1 Answers
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Modeling Chemistry Unit 7 1 Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Modeling Chemistry Unit 7 1 Answers
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Modeling Chemistry Unit 7 1 Answers
  - Personalized Recommendations
  - Modeling Chemistry Unit 7 1 Answers User Reviews and Ratings
  - Modeling Chemistry Unit 7 1 Answers and Bestseller Lists

- 5. Accessing Modeling Chemistry Unit 7 1 Answers Free and Paid eBooks
  - Modeling Chemistry Unit 7 1 Answers Public Domain eBooks
  - Modeling Chemistry Unit 7 1 Answers eBook Subscription Services
  - Modeling Chemistry Unit 7 1 Answers Budget-Friendly Options
- 6. Navigating Modeling Chemistry Unit 7 1 Answers eBook Formats
  - o ePub, PDF, MOBI, and More
  - Modeling Chemistry Unit 7 1 Answers Compatibility with Devices
  - Modeling Chemistry Unit 7 1 Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Modeling Chemistry Unit 7 1 Answers
  - Highlighting and Note-Taking Modeling Chemistry Unit 7 1 Answers
  - Interactive Elements Modeling Chemistry Unit 7 1 Answers
- 8. Staying Engaged with Modeling Chemistry Unit 7 1 Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Modeling Chemistry Unit 7 1 Answers
- 9. Balancing eBooks and Physical Books Modeling Chemistry Unit 7 1 Answers
  - ∘ Benefits of a Digital Library
  - $\circ$  Creating a Diverse Reading Collection Modeling Chemistry Unit 7 1 Answers
- 10. Overcoming Reading Challenges
  - o Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Modeling Chemistry Unit 7 1 Answers
  - Setting Reading Goals Modeling Chemistry Unit 7 1 Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Modeling Chemistry Unit 7 1 Answers
  - Fact-Checking eBook Content of Modeling Chemistry Unit 7 1 Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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