

## Thermodynamics Review

1. What is a spontaneous change or reaction?

Spontaneous changes or reactions occur in a forward direction without any outside intervention.

2. Will a spontaneous reaction *always* proceed at a rapid rate towards equilibrium?

No, it will if given enough time, but thermodynamics cannot predict how fast the reaction will proceed. Kinetics is the study of reaction rates.

3. Because a reaction is spontaneous, does it mean that a lot of product will be produced when the system reaches equilibrium?

The spontaneity of a reaction does not tell you whether a lot of product will be produced or not. (A reaction that has a large negative value of  $\Delta G$ , however will usually be product favored.)

We will talk more about equilibrium in Chapter 15.

4. A chemical reaction in which most of the reactants can eventually be converted to products, *if given sufficient time*, is said to be product favored.

5. If a reaction is reactant favored, does it mean that it will not occur at all?

No, it just means that not a lot of product is formed.

Chemists spend a lot of time and money varying reaction conditions to maximize the amount of product formed.

We will talk about this more next unit.

6. The thermodynamic function called entropy is a measure of the matter dispersal or disorder of the system.

7. All spontaneous processes increase the total entropy of the universe (2<sup>nd</sup> Law)

8. If the entropy of each element is taken as zero at 0 K, every substance has a finite entropy. This statement is called the 3<sup>rd</sup> Law of Thermodynamics.

### PREDICT S

9. By inspection, predict whether the entropy value will increase or decrease during the following processes and give the sign of the entropy change:



### CALCULATE S

# Modeling Chemistry Unit 6 Review Answers

**M Lipman**



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